



UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

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CHEMISTRY

0620/11

Paper 1 Multiple Choice

May/June 2013

45 Minutes

Additional Materials: Multiple Choice Answer Sheet
Soft clean eraser
Soft pencil (type B or HB is recommended)

* 9 6 3 9 7 8 9 7 1 6 *

READ THESE INSTRUCTIONS FIRST

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

DO NOT WRITE IN ANY BARCODES.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A, B, C** and **D**.

Choose the **one** you consider correct and record your choice in **soft pencil** on the separate Answer Sheet.

Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.

Electronic calculators may be used.

This document consists of **15** printed pages and **1** blank page.



- 1 The diagram shows a cup of tea.



Which row describes the water particles in the air above the cup compared with the water particles in the cup?

	moving faster	closer together
A	✓	✓
B	✓	x
C	x	✓
D	x	x

- 2 Crystals of sodium chloride were prepared by the following method.

- 1 25.0 cm³ of dilute hydrochloric acid was accurately measured into a conical flask.
- 2 Aqueous sodium hydroxide was added until the solution was neutral. The volume of sodium hydroxide added was measured.
- 3 The solution was evaporated and the crystals washed with approximately 15 cm³ of water.

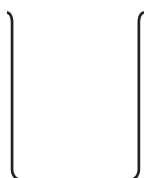
Which row shows the pieces of apparatus used to measure the 25.0 cm³ of hydrochloric acid, the volume of aqueous sodium hydroxide and the 15 cm³ of water?

	25.0 cm ³ of hydrochloric acid accurately	the volume of aqueous sodium hydroxide added	15 cm ³ of water approximately
A	burette	pipette	measuring cylinder
B	measuring cylinder	burette	pipette
C	pipette	burette	measuring cylinder
D	pipette	measuring cylinder	burette

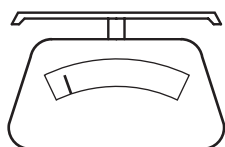
- 3 Lead iodide is insoluble in water.

Lead iodide is made by adding aqueous lead nitrate to aqueous potassium iodide.

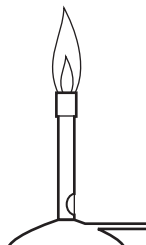
Which pieces of apparatus are needed to obtain solid lead iodide from 20 cm³ of aqueous lead nitrate?



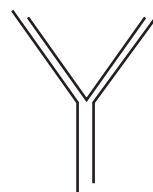
1



2



3



4



5

A 1, 2 and 4

B 1, 3 and 5

C 1, 4 and 5

D 2, 4 and 5

- 4 Element X is represented by $^{27}_{13}\text{X}$.

Which statement about element X is correct?

- A An atom of X contains 13 protons and 13 neutrons.
 B An atom of X contains 27 protons and 13 electrons.
 C X forms an ion by gaining electrons.
 D X is placed in Group III of the Periodic Table.

- 5 The positions of four elements are shown on the outline of the Periodic Table.

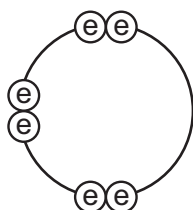
Which element forms a coloured oxide?

																					A		
	B																						C
							D																

6 For which substance is the type of bonding **not** correct?

	substance	type of bonding		
		ionic	covalent	metallic
A	chlorine		✓	
B	potassium bromide	✓		
C	sodium			✓
D	sodium chloride		✓	

7 Element X has six electrons in its outer shell.

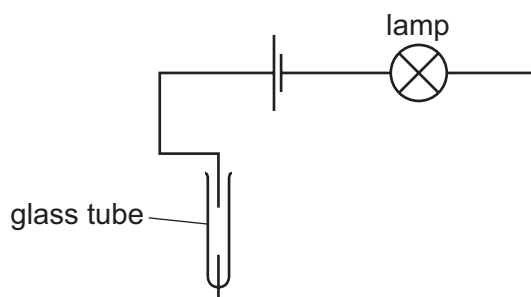


key

⊕ = electron

How could the element react?

- A** by gaining two electrons to form a positive ion
 - B** by losing six electrons to form a negative ion
 - C** by sharing two electrons with two electrons from another element to form two covalent bonds
 - D** by sharing two electrons with two electrons from another element to form four covalent bonds
- 8 The diagram shows an incomplete circuit.



Which substance causes the lamp to light when added to the glass tube?

- A** aqueous sodium chloride
- B** aqueous sugar
- C** solid sodium chloride
- D** solid sugar

9 A compound with the formula XF_2 has a relative formula mass of 78.

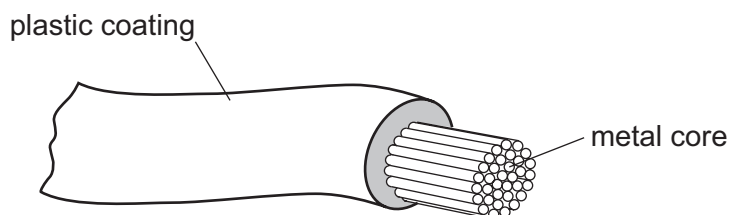
What is element X?

- A argon
- B calcium
- C neon
- D zirconium

10 What is the balanced chemical equation for the reaction between calcium and water?

- A $\text{Ca} + \text{H}_2\text{O} \rightarrow \text{CaOH} + \text{H}_2$
- B $\text{Ca} + \text{H}_2\text{O} \rightarrow \text{Ca}(\text{OH})_2 + \text{H}_2$
- C $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{CaOH} + \text{H}_2$
- D $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca}(\text{OH})_2 + \text{H}_2$

11 The diagram shows an electrical cable.

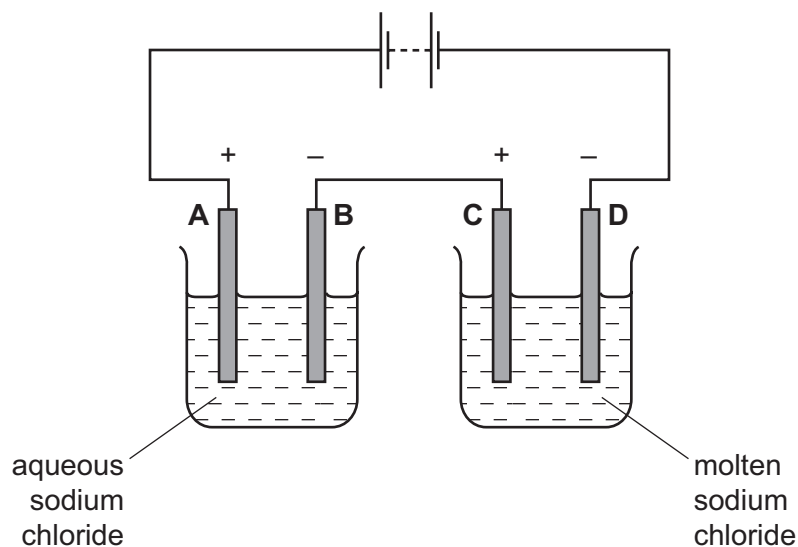


Which statement about the substances used is correct?

- A The coating is plastic because it conducts electricity well.
- B The core is copper because it conducts electricity well.
- C The core is copper because it is cheap and strong.
- D The core is iron because it is cheap and strong.

12 The diagram shows an electrolysis circuit.

At which electrode is hydrogen formed?



13 Some white anhydrous copper(II) sulfate powder is put into a beaker of water and stirred.

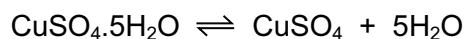
What would show that the process was exothermic?

- A A blue solution is formed.
- B The beaker feels cooler.
- C The beaker feels warmer.
- D The powder dissolves in the water.

14 Which substance does **not** require oxygen in order to produce energy?

- A coal
- B hydrogen
- C natural gas
- D ^{235}U

- 15 The equation shows the formation of anhydrous copper(II) sulfate from hydrated sulfate.



Statements 1, 2 and 3 refer to this reaction.

- 1 Hydrated copper(II) sulfate is reduced to anhydrous copper(II) sulfate.
- 2 The (II) in the name copper(II) sulfate refers to the oxidation state of the metal.
- 3 The reaction is reversible.

Which statements are correct?

- A** 1 only **B** 1 and 2 **C** 2 and 3 **D** 3 only

- 16 Calcium carbonate reacts with hydrochloric acid to form carbon dioxide.

Which changes would slow this reaction down?

- 1 decreasing the concentration of hydrochloric acid
- 2 decreasing the particle size of calcium carbonate
- 3 decreasing the temperature

- A** 1 and 2 only **B** 1 and 3 only **C** 2 and 3 only **D** 1, 2 and 3

- 17 The equations represent redox reactions.

In which equation is the underlined substance acting as a reducing agent?

- A** 3CO + Fe₂O₃ → 2Fe + 3CO₂
- B** CO₂ + C → 2CO
- C** CuO + H₂ → Cu + H₂O
- D** CaO + H₂O → Ca(OH)₂

- 18 Ant stings hurt because of the methanoic acid produced by the ant.

Which substance could, **most safely**, be used to neutralise the acid?

	substance	pH
A	baking soda	8
B	car battery acid	1
C	lemon juice	3
D	oven cleaner	14

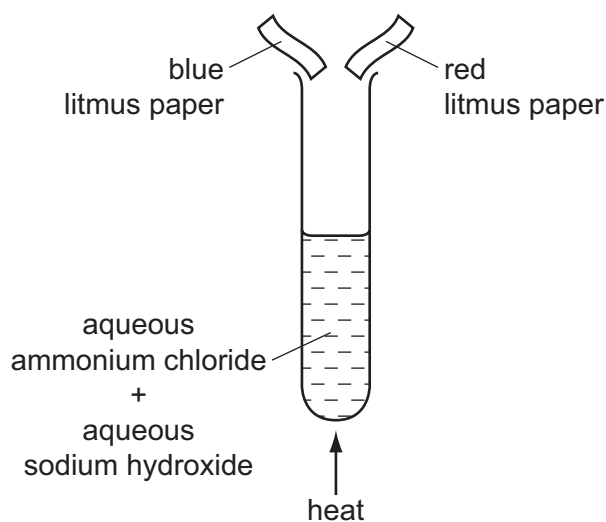
19 The diagram shows one period of the Periodic Table.

Li	Be	B	C	N	O	F	Ne
----	----	---	---	---	---	---	----

Which two elements form acidic oxides?

- A carbon and lithium
- B carbon and neon
- C carbon and nitrogen
- D nitrogen and neon

20 The diagram shows an experiment.



What happens to the pieces of litmus paper?

	blue litmus paper	red litmus paper
A	changes colour	changes colour
B	changes colour	no colour change
C	no colour change	changes colour
D	no colour change	no colour change

- 21 Two indicators, bromophenol blue and Congo red, show the following colours in acid and in alkaline solutions.

indicator	acid	alkali
bromophenol blue	yellow	blue
Congo red	violet	red

A few drops of each indicator are added to separate samples of a solution of pH 2.

What are the colours of the indicators in this solution?

	in a solution of pH 2	
	bromophenol blue is	Congo red is
A	blue	red
B	blue	violet
C	yellow	red
D	yellow	violet

- 22 Which property of elements increases across a period of the Periodic Table?

- A** metallic character
- B** number of electron shells
- C** number of outer shell electrons
- D** tendency to form positive ions

- 23 Which element is a transition metal?

	colour of chloride	melting point of element / °C
A	white	113
B	white	1495
C	yellow	113
D	yellow	1495

24 Fluorine is at the top of Group VII in the Periodic Table.

Which row shows the properties of fluorine?

	colour	state at room temperature	reaction with aqueous potassium iodide
A	brown	gas	no reaction
B	brown	liquid	iodine displaced
C	yellow	gas	iodine displaced
D	yellow	liquid	no reaction

25 Group I metals are also known as the Alkali Metals.

Which statement about the metals in Group I is **not** correct?

- A** In their reactions they lose electrons.
- B** Their atoms all have one electron in their outer shell.
- C** They form +1 ions in their reactions with non-metals.
- D** They form covalent compounds by sharing electrons.

26 Which element is a metal?

	charge on element ion	electrical conductivity
A	negative	low
B	positive	high
C	negative	high
D	positive	low

27 Which property makes aluminium ideal for making food containers?

- A** conducts electricity
- B** conducts heat
- C** mechanical strength
- D** resistance to corrosion

28 Which substance is **not** involved in the extraction of iron from hematite?

- A carbon
- B carbon monoxide
- C calcium carbonate
- D nitrogen

29 Pure metals conduct electricity and can be hammered into different shapes.

Why are metals sometimes used as alloys?

- A Alloys are cheaper than the metals they are made from.
- B Alloys are easier to hammer into different shapes.
- C Alloys are harder and keep their shape better.
- D Alloys conduct electricity better.

30 Below are some metals in decreasing order of reactivity.

magnesium

zinc

iron

copper

Titanium reacts with acid and cannot be extracted from its ore by heating with carbon.

Where should titanium be placed in this list?

- A below copper
- B between iron and copper
- C between magnesium and zinc
- D between zinc and iron

31 Water has been contaminated with sea-water.

Which substances can be removed by chlorination and filtration?

- A bacteria, sand and sodium chloride
- B bacteria and sand only
- C bacteria and sodium chloride only
- D sand and sodium chloride only

32 Iron rusts when it reacts with1.....

Rusting can be prevented by covering the iron with a more reactive metal, such as2.....

Which words correctly complete gaps 1 and 2?

	1	2
A	oxygen	copper
B	oxygen	magnesium
C	oxygen and water	copper
D	oxygen and water	magnesium

33 Nitrogen, phosphorus and potassium are essential elements for plant growth.

Which mixture provides all three essential elements?

	mixture	formula
A	ammonium phosphate + potassium chloride	$(\text{NH}_4)_3\text{PO}_4$ + KCl
B	ammonium phosphate + ammonium nitrate	$(\text{NH}_4)_3\text{PO}_4$ + NH_4NO_3
C	ammonium phosphate + ammonium chloride	$(\text{NH}_4)_3\text{PO}_4$ + NH_4Cl
D	ammonium nitrate + potassium chloride	NH_4NO_3 + KCl

34 Which information about carbon dioxide and methane is correct?

		carbon dioxide	methane
A	formed when vegetation decomposes	✓	✗
B	greenhouse gas	✓	✓
C	present in unpolluted air	✗	✗
D	produced during respiration	✗	✓

key
✓ = true
✗ = false

35 The list shows four methods that were suggested for the formation of carbon dioxide.

- 1 action of an alkali on a carbonate
- 2 action of heat on a carbonate
- 3 complete combustion of methane
- 4 reaction of a carbonate with oxygen

Which methods would result in the production of carbon dioxide?

- A** 1 and 2 **B** 1 and 4 **C** 2 and 3 **D** 3 and 4

36 Organic compounds may have names ending in -ane, -ene, -ol or -oic acid.

How many of these endings indicate the compounds contain double bonds in their molecules?

- A** 1 **B** 2 **C** 3 **D** 4

37 The table shows the boiling points of four members of the homologous series of alcohols.

compound		boiling point /°C
name	formula	
methanol	CH ₃ OH	65
ethanol	C ₂ H ₅ OH	78
propanol	C ₃ H ₇ OH	X
butanol	C ₄ H ₉ OH	117

What is the value of X?

- A** 55°C **B** 82°C **C** 98°C **D** 115°C

- 38 The table shows some fractions that are obtained from petroleum by fractional distillation together with some of their uses.

fraction	use
refinery gas	cooking
gasoline	fuel for cars
1	making chemicals
2	jet fuel
3	fuel for ships
bitumen	making roads

Which row correctly identifies fractions 1, 2 and 3?

	1	2	3
A	diesel oil	fuel oil	lubricating fraction
B	fuel oil	diesel oil	kerosene
C	kerosene	naphtha	diesel oil
D	naphtha	kerosene	fuel oil

- 39 Which columns describe the hydrocarbons ethane and ethene?

	1	2	3	4
state at room temperature	gas	gas	liquid	liquid
reaction with oxygen	burns	burns	burns	burns
reaction with aqueous bromine	no reaction	decolourises bromine	no reaction	decolourises bromine

- A** 1 (ethane) and 2 (ethene)
B 1 (ethane) and 4 (ethene)
C 2 (ethene) and 3 (ethane)
D 3 (ethane) and 4 (ethene)
- 40 Which of the statements about ethanol are correct?
- Ethanol can be formed by an addition reaction.
 - Ethanol can be formed by fermentation.
 - When ethanol burns in air, it forms carbon dioxide and water.
- A** 1, 2 and 3 **B** 1 and 2 **C** 1 and 3 **D** 2 and 3

DATA SHEET The Periodic Table of the Elements

		Group																							
		I	II	III	IV	V	VI	VII	0																
7	9	1									4														
Li Lithium 3	Be Beryllium 4	H Hydrogen 1									He Helium 2														
23	24																								
Na Sodium 11	Mg Magnesium 12																								
39	40	51	48	45	91	93	96	55	56	59	64	65	70	73	75	79	80	84							
K Potassium 19	Ca Calcium 20	V Vanadium 23	Ti Titanium 22	Sc Scandium 21	Zr Zirconium 40	Nb Niobium 41	Cr Chromium 24	Mn Manganese 25	Fe Iron 26	Co Cobalt 27	Ni Nickel 28	Cu Copper 29	Zn Zinc 30	Ga Gallium 31	Ge Germanium 32	As Arsenic 33	Se Selenium 34	Br Bromine 35	Kr Krypton 36						
85	88	137	178	89	91	93	96	101	106	112	127	131	133	139	141	144	149	150	157	162	167	173	175		
Rb Rubidium 37	Sr Strontium 38	Cs Caesium 55	Ba Barium 56	Y Yttrium 39	Zr Zirconium 40	Nb Niobium 41	Mo Molybdenum 42	Ru Ruthenium 44	Rh Rhodium 45	Pd Palladium 46	Ag Silver 47	Cd Cadmium 48	In Indium 49	Sn Tin 50	Sb Antimony 51	Te Tellurium 52	I Iodine 53	Xe Xenon 54	209	207	210	211	212	214	217
Fr Francium 87	Ra Radium 88											226	227	227	227	227	227	227	227	227	227	227	227		

140	141	144	150	152	157	159	162	165	167	169	173	175
Ce Cerium 58	Pr Praseodymium 59	Nd Neodymium 60	Pm Promethium 61	Sm Samarium 62	Eu Europium 63	Gd Gadolinium 64	Dy Dysprosium 66	Ho Holmium 67	Er Erbium 68	Tm Thulium 69	Yb Ytterbium 70	Lu Lutetium 71
232	238	238	238	238	238	238	238	238	238	238	238	238
Th Thorium 90	Pa Protactinium 91	U Uranium 92	Np Neptunium 93	Pu Plutonium 94	Am Americium 95	Cm Curium 96	Bk Berkelium 97	Cf Californium 98	Es Einsteinium 99	Fm Fermium 100	Md Mendelevium 101	Lr Lawrencium 103

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

Key	a	X	a = relative atomic mass
	b	X	X = atomic symbol b = proton (atomic) number

*58-71 Lanthanoid series
†90-103 Actinoid series

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