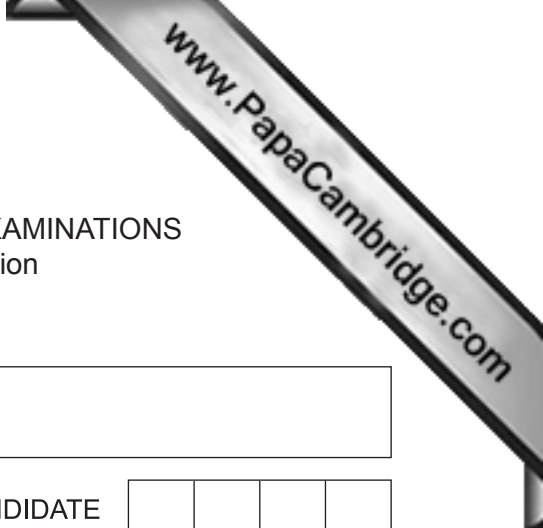




UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education



CANDIDATE
NAME

CENTRE
NUMBER

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CANDIDATE
NUMBER

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CHEMISTRY

0620/22

Paper 2

October/November 2013

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name in the spaces at the top of this page.

Write in dark blue or black pen.

You may need to use a pencil for any diagrams, graphs or rough working.

Do not use staples, paper clips, highlighters, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 16.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

This document consists of **15** printed pages and **1** blank page.



1 (a) Choose from the list of compounds below to answer the following questions.

- ammonia
- ammonium chloride
- calcium carbonate
- calcium oxide
- copper(II) sulfate
- ethane
- iron(II) chloride
- methane
- water

Each compound can be used once, more than once or not at all.

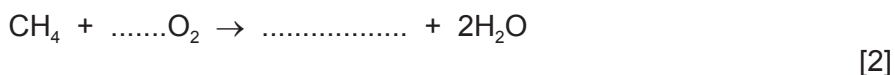
Which compound:

- (i) is an alkaline gas, [1]
- (ii) is a gas contributing to climate change, [1]
- (iii) is a salt containing only non-metals, [1]
- (iv) turns blue cobalt chloride paper pink, [1]
- (v) reacts with an acid to release carbon dioxide, [1]
- (vi) gives a light blue precipitate when aqueous sodium hydroxide is added to a solution of its aqueous ions? [1]

(b) What is the meaning of the term *compound*?

.....
..... [1]

(c) Complete the following symbol equation for the complete combustion of methane in oxygen.



[Total: 9]

- 2 (a) The table describes the reactivity of some metals with hydrochloric acid.

metal	observations
calcium	Many bubbles produced. Reaction mixture may boil.
magnesium	Steady stream of bubbles produced. Reaction mixture gets hot.
sodium	Many bubbles produced. May explode.
zinc	Slow stream of bubbles produced. Reaction mixture rises slightly in temperature.

Put these metals in order of their reactivity.

least reactive \longrightarrow most reactive

--	--	--	--

[2]

- (b) Complete the word equation for the reaction of magnesium with hydrochloric acid.

magnesium + hydrochloric acid \rightarrow +

.....

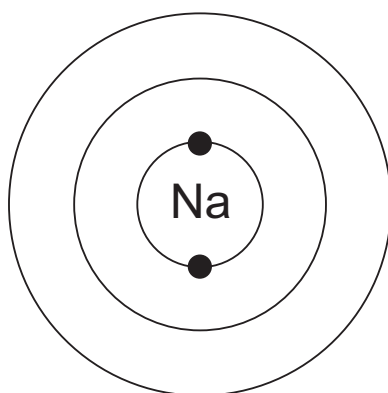
[2]

- (c) When magnesium reacts with hydrochloric acid, magnesium atoms lose electrons. What type of magnesium particle is formed? Put a ring around the correct answer.

covalent **ion** **molecule** **proton**

[1]

- (d) Complete the diagram to show the electronic structure of a sodium atom.



[2]

(e) A student added large lumps of zinc to 20 cm³ of 2 mol / dm³ hydrochloric acid. She carried out the reaction at 15 °C. She measured the volume of gas given off at various time intervals.

(i) Draw a labelled diagram of the apparatus she could use for this experiment.

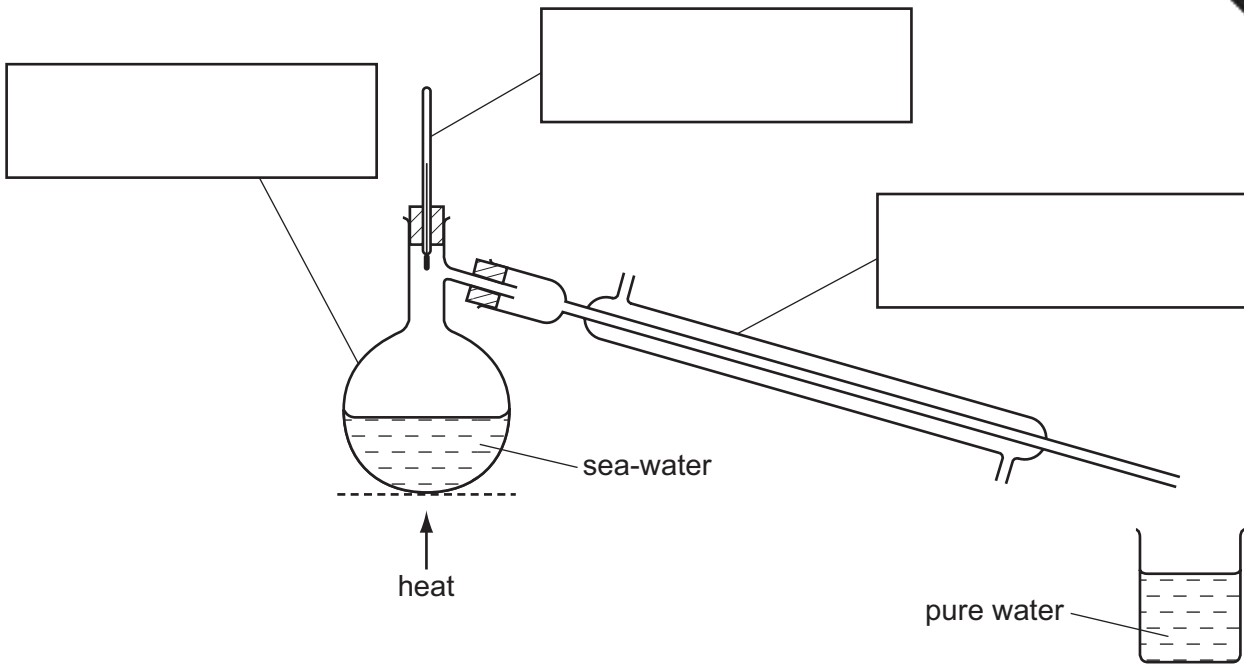
[3]

(ii) Describe **three** different things she could do to increase the rate of this reaction.

- 1.
- 2.
- 3. [3]

[Total: 13]

3 The diagram below shows the apparatus which can be used to obtain pure water from sea-water.



(a) State the name of this process.

..... [1]

(b) Label the boxes on the diagram above with the correct names of the pieces of apparatus shown. [3]

(c) Complete the following sentences using words from the list below.

- | | | | |
|---------------|------------------|--------------|----------------|
| boils | condenses | cools | freezes |
| higher | lower | melts | |

Water has a boiling point than salt. When a solution of salt is heated strongly, the water and escapes as steam. When the steam cools, it back to liquid water. [3]

(d) The table shows the concentration of the seven most abundant compounds in sea water.

compound	ions present	concentration in g/m ³
calcium carbonate	Ca ²⁺ and CO ₃ ²⁻	100
calcium sulfate	Ca ²⁺ and SO ₄ ²⁻	1 800
magnesium chloride	Mg ²⁺ and Cl ⁻	6 800
magnesium sulfate		5 700
potassium bromide	K ⁺ and Br ⁻	100
potassium chloride	K ⁺ and Cl ⁻	800
sodium chloride	Na ⁺ and Cl ⁻	28 000

(i) Which negative ion is present in the greatest concentration in sea-water?

..... [1]

(ii) Which positive ion is present in the lowest concentration in sea-water?

..... [1]

(iii) Write the formulae of the **two** ions present in magnesium sulfate.

..... [2]

[Total: 11]

- 4 (a) Match the compounds on the left with the statements on the right.
The first one has been done for you.

butane	a hydrocarbon containing four carbon atoms
poly(ethene)	it decolourises bromine water
ethene	it is the main constituent of natural gas
methane	it contains a -COOH functional group
ethanoic acid	it has a very long chain of carbon atoms

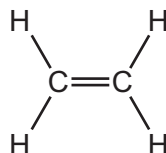
[4]

- (b) Methane and ethene are hydrocarbons.

- (i) What is meant by the term *hydrocarbon*?

..... [1]

- (ii) The structure of ethene is shown below.



Use this structure to explain why ethene is an unsaturated hydrocarbon.

..... [1]

- (c) Molecules of ethene react together at high temperature and pressure to form poly(ethene).

Which **one** of the following words best describes the molecules of ethene in this reaction?
Put a ring around the correct answer.

acids alkanes monomers polymers

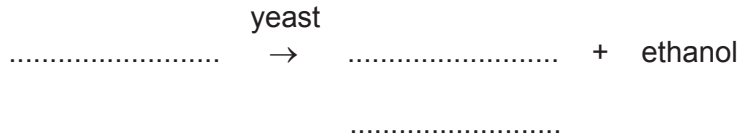
[1]

(d) Ethanoic acid can be made by the oxidation of ethanol.

(i) What is meant by the term *oxidation*?

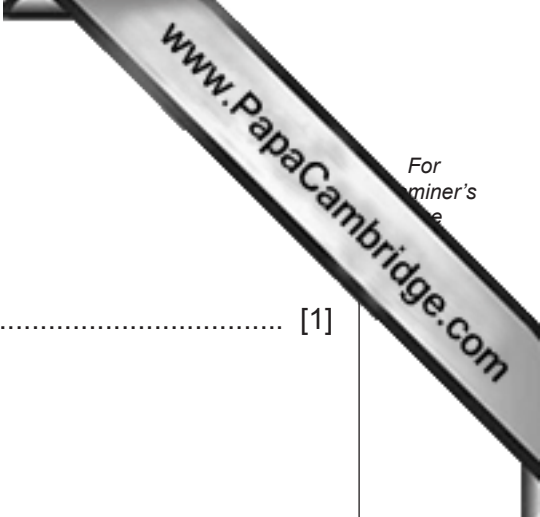
..... [1]

(ii) Ethanol can be made by fermentation.
Complete the word equation for fermentation.



[2]

[Total: 10]



For
miner's
e

5 (a) Explain why metals are often used in the form of alloys.
In your answer, write about

- the structure of an alloy,
- why alloys are often more useful than pure metals.

.....

.....

.....

.....

..... [3]

(b) Iron is a transition element.

(i) Which two of the following statements about iron are correct?
Tick **two** boxes.

A freshly-cut surface of iron is green in colour.

Iron exists in only one oxidation state in its compounds.

Iron has a high density.

Iron has a giant covalent structure.

Iron has a high melting point.

[2]

(ii) Describe **one** method of rust prevention and explain how it works.

method

how this works

..... [2]

(c) Iron is used as a catalyst in the Haber process for making ammonia.

(i) What does the term *catalyst* mean?

..... [1]

(ii) Describe a test for ammonia.

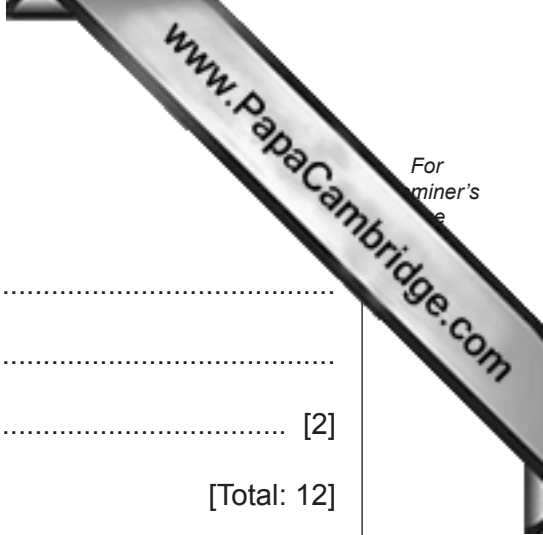
test

result [2]

- (iii) Ammonia is used to make fertilisers.
Explain why farmers need to add fertilisers to the soil.

.....
.....
..... [2]

[Total: 12]



For
miner's
e

- 6 (a) Garlic is a vegetable that is often used in cooking. It has a strong smell. A student is cutting up garlic in the kitchen.



After a time, the smell of the garlic travels all over the house even though there are no currents of air.

Use the kinetic particle theory to explain why the smell of garlic travels all over the house.

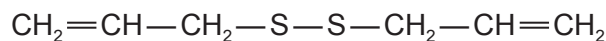
.....

.....

.....

..... [3]

- (b) The smell of garlic is due to a compound containing sulfur. The structure of this compound (compound **A**) is shown below.

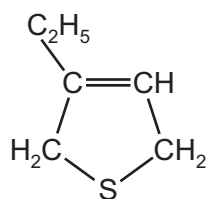


compound **A**

- (i) Write the molecular formula for this compound.

..... [1]

- (ii) Another organic sulfur compound (compound **B**) is shown below.



compound **B**

By comparing the formulae of compound **A** and compound **B**, how can you tell that compound **A** has the higher relative molecular mass?

You are not required to do any mathematical calculations.

.....

..... [2]

(c) An isotope of sulfur has a nucleon number of 34 and an atomic number of 16.

(i) How many neutrons are there in one atom of this isotope of sulfur?

..... [1]

(ii) What is meant by the terms

isotope,

..... [1]

nucleon number? [1]

(iii) Some fuels contain sulfur as a contaminating substance.
Complete the following sentences using words from the list below.

- | | | | |
|-----------------|-----------------|-----------------|-----------------|
| coal | dioxide | hydrogen | monoxide |
| nitrogen | oxidised | reduced | water |

Fuels such as contain sulfur.

When these fuels burn, the sulfur is to sulfur

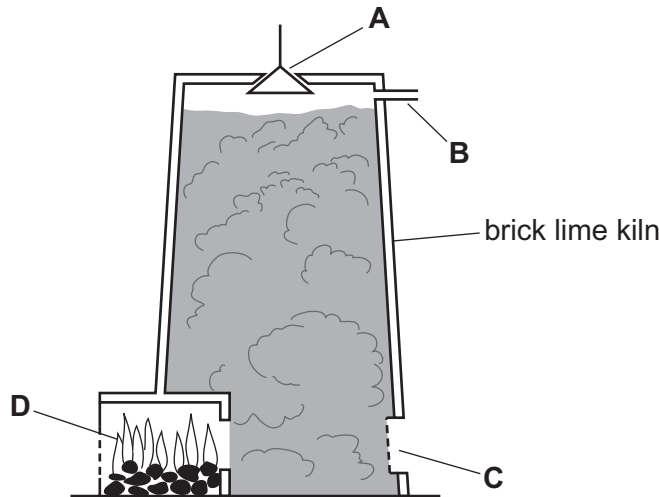
This reacts with in the atmosphere to form an acidic solution. [4]

(iv) Describe and explain the effect of acid rain on buildings made of limestone.

.....
.....
..... [2]

[Total: 15]

7 The diagram shows a kiln for making lime (calcium oxide) from limestone (calcium carb



- (a) (i) Which letter on the diagram above shows
 where the limestone is added,
 where the waste gases exit from the kiln? [2]

(ii) Complete the symbol equation for the decomposition of limestone.



(iii) When 50 g of calcium carbonate is decomposed, 28 g of calcium oxide is formed.
 Calculate the minimum mass of calcium carbonate needed to produce 8.4 g of
 calcium oxide.

[1]

(b) The table below shows the temperatures at which some Group II carbonates decompose.

Group II carbonate	temperature at which Group II carbonates decompose / °C
beryllium carbonate	100
magnesium carbonate	350
calcium carbonate	900

- (i) Describe the pattern in the ease of decomposition of Group II carbonates.
 [1]

(ii) Predict the decomposition temperature of barium carbonate.

..... °C [1]

(c) Lime is calcium oxide.

(i) State **one** use of lime.

..... [1]

(ii) What type of oxide is calcium oxide?

..... [1]

(iii) Calculate the relative formula mass of calcium oxide.
Use your Periodic Table to help you.

[1]

(d) Calcium is extracted from its compounds by electrolysis.
Suggest why calcium is extracted by electrolysis rather than by reduction with carbon.

..... [1]

[Total: 10]

DATA SHEET
The Periodic Table of the Elements

		Group																																																																		
I	II	III	IV	V	VI	VII	0																																																													
7 Li Lithium 3	9 Be Beryllium 4	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	13 Al Aluminium 13	14 N Nitrogen 7	15 O Oxygen 8	16 F Fluorine 9	17 Ne Neon 10	18 Ar Argon 18	19 K Potassium 19	20 Ca Calcium 20	21 Sc Scandium 21	22 Ti Titanium 22	23 V Vanadium 23	24 Cr Chromium 24	25 Mn Manganese 25	26 Fe Iron 26	27 Co Cobalt 27	28 Ni Nickel 28	29 Cu Copper 29	30 Zn Zinc 30	31 Ga Gallium 31	32 Ge Germanium 32	33 As Arsenic 33	34 Se Selenium 34	35 Br Bromine 35	36 Kr Krypton 36	37 Rb Rubidium 37	38 Sr Strontium 38	39 Y Yttrium 39	40 Zr Zirconium 40	41 Nb Niobium 41	42 Mo Molybdenum 42	43 Tc Technetium 43	44 Ru Ruthenium 44	45 Rh Rhodium 45	46 Pd Palladium 46	47 Ag Silver 47	48 Cd Cadmium 48	49 In Indium 49	50 Sn Tin 50	51 Sb Antimony 51	52 Te Tellurium 52	53 I Iodine 53	54 Xe Xenon 54	55 Cs Caesium 55	56 Ba Barium 56	57 La Lanthanum 57	72 Hf Hafnium 72	73 Ta Tantalum 73	74 W Tungsten 74	75 Re Rhenium 75	76 Os Osmium 76	77 Ir Iridium 77	78 Pt Platinum 78	79 Au Gold 79	80 Hg Mercury 80	81 Tl Thallium 81	82 Pb Lead 82	83 Bi Bismuth 83	84 Po Polonium 84	85 At Astatine 85	86 Rn Radon 86	87 Fr Francium 87	88 Ra Radium 88	89 Ac Actinium 89	†
												133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	178 Hf Hafnium 72	181 Ta Tantalum 73	184 W Tungsten 74	186 Re Rhenium 75	190 Os Osmium 76	192 Ir Iridium 77	195 Pt Platinum 78	197 Au Gold 79	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	212 Po Polonium 84	210 At Astatine 85	210 Rn Radon 86	226 Fr Francium 87	227 Ra Radium 88	227 Ac Actinium 89	†																																			
												140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	159 Tb Terbium 65	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	232 Th Thorium 90	238 U Uranium 92	238 Pa Protactinium 91	238 Np Neptunium 93	238 Pu Plutonium 94	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102	238 Lr Lawrencium 103																														

*58-71 Lanthanoid series
†90-103 Actinoid series

Key

a	X
b	

 a = relative atomic mass
 X = atomic symbol
 b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).

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