

CANDIDATE
NAME

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CHEMISTRY

0620/32

Paper 3 (Extended)

October/November 2015

1 hour 15 minutes

Candidates answer on the Question Paper.

No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO NOT WRITE IN ANY BARCODES.

Answer **all** questions.

Electronic calculators may be used.

A copy of the Periodic Table is printed on page 12.

You may lose marks if you do not show your working or if you do not use appropriate units.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

The syllabus is approved for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.

This document consists of **12** printed pages.

1 Use your copy of the Periodic Table to help you answer some of these questions.

(a) Predict the formulae of the following compounds.

(i) nitrogen fluoride

(ii) phosphorus sulfide

[2]

(b) Deduce the formulae of the following ions.

(i) selenide

(ii) gallium

[2]

(c) Use the following ions to determine the formulae of the compounds.

ions OH^- Cr^{3+} Ba^{2+} SO_4^{2-}

compounds

(i) chromium(III) sulfate

(ii) barium hydroxide

[2]

[Total: 6]

2 (a) Polluted air contains two oxides of carbon and two oxides of nitrogen. A major source of these pollutants is motor vehicles.

(i) Describe how carbon dioxide and carbon monoxide are formed in motor vehicle engines.

.....
.....
.....
..... [3]

(ii) State **one** adverse effect of each of these gases.

.....
..... [2]

(iii) Nitrogen monoxide, NO, is released by motor vehicle exhausts.

Explain how nitrogen monoxide is formed in motor vehicle engines.

.....
..... [2]

(iv) When nitrogen monoxide is released into the atmosphere, nitrogen dioxide, NO₂, is formed.

Suggest an explanation why this happens.

..... [1]

(b) Predict the possible adverse effect on the environment when this non-metal oxide, NO₂, reacts with water and oxygen.

.....
..... [2]

(c) How are the amounts of carbon monoxide and nitrogen monoxide emitted by modern motor vehicles reduced? Include an equation in your answer.

.....
.....
..... [3]

[Total: 13]

- 3 Two of the main uses of zinc are for galvanising and for making alloys.

One of the main ores of zinc is zinc blende, ZnS. There are two stages in the extraction of zinc from this ore.

- (a) **Stage 1** Zinc oxide is made from zinc blende.

Describe how this is done and write a word equation for the reaction.

.....

 [2]

- (b) **Stage 2** Zinc oxide is reduced to zinc.

Write a word equation for the reduction of zinc oxide by coke.

..... [1]

- (c) The zinc produced by this process is impure. It can be purified by electrolysis using a method which is similar to the purification of copper. Under the conditions used in the process, zinc is the product at the negative electrode (cathode).

Complete the following description of this purification.

The electrolyte is aqueous [1]

The negative electrode (cathode) is made of [1]

The positive electrode (anode) is impure zinc.

The equation for the reaction at the cathode is [1]

The equation for the reaction at the anode is [1]

Explain why the concentration of the electrolyte does **not** change.

.....
 [2]

(d) Brass is an alloy which contains zinc.

(i) Name the other metal in brass.

..... [1]

(ii) Suggest **two** reasons why an alloy such as brass is preferred to either of its constituent metals.

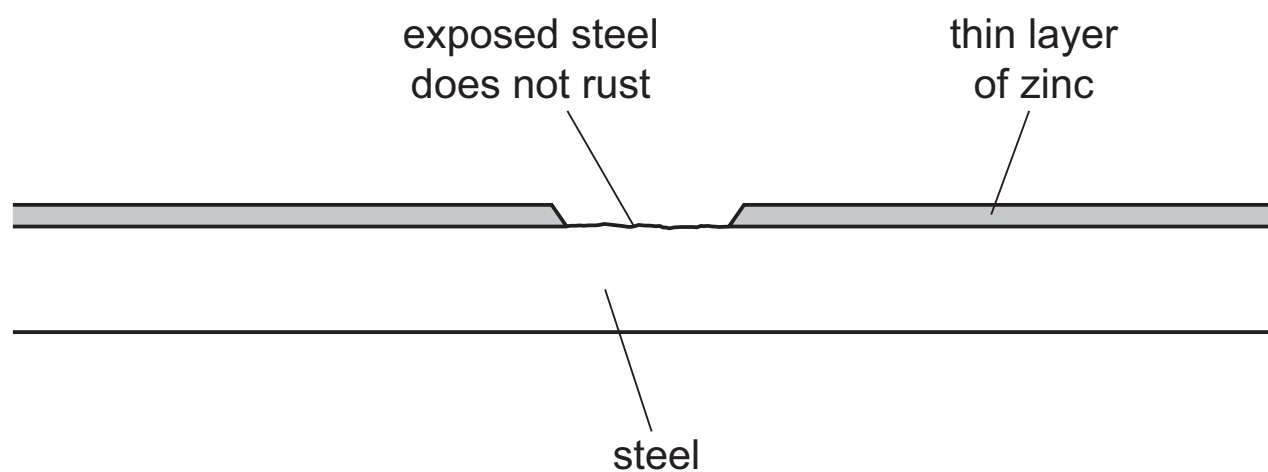
.....
 [2]

(e) In an experiment to investigate the rate of rusting of steel, three pieces of steel were used. One piece of steel was completely coated with copper, one piece completely coated with zinc and the third piece was left uncoated. All three pieces were left exposed to the atmosphere.

(i) Explain why the uncoated piece started to rust.

.....
 [1]

(ii) The coating on both of the other two pieces was scratched, exposing the steel.



The piece of steel coated with zinc still did not rust but the copper-coated piece of steel rusted very rapidly.

Explain these observations in terms of the formation of ions and the transfer of electrons.

.....

 [4]

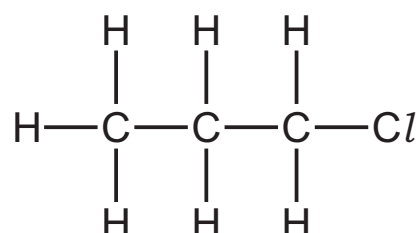
[Total: 17]

4 (a) Propane reacts with chlorine to form a mixture of chloropropanes. This is a photochemical reaction.

(i) What is meant by the phrase *photochemical reaction*?

.....
 [1]

(ii) The products of this reaction include two isomers, one of which has the following structural formula.



Draw the structural formula of the other isomer.

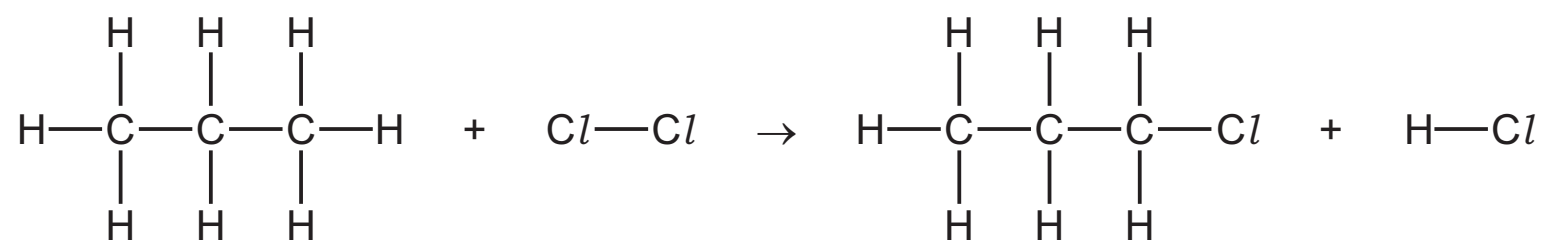
[1]

(iii) Explain why these two different compounds are isomers.

.....
 [2]

(b) Bond breaking is an endothermic change and bond forming is an exothermic change.

Bond energy is the amount of energy in kJ/mol needed to break one mole of the specified bond.



Use the following bond energies to determine whether this reaction is exothermic or endothermic. You must show your reasoning.

bond	bond energies in kJ/mol
C-Cl	338
C-H	412
Cl-Cl	242
H-Cl	431
C-C	348

.....

.....

.....

..... [3]

- (c) (i) Chloropropane can be hydrolysed to propanol, $\text{CH}_3\text{CH}_2\text{CH}_2\text{OH}$, by sodium hydroxide.

Write the equation for this reaction.

..... [2]

- (ii) Propanol can be dehydrated. It loses a water molecule to form a hydrocarbon.

Give the name and structural formula of this hydrocarbon.

name

structural formula

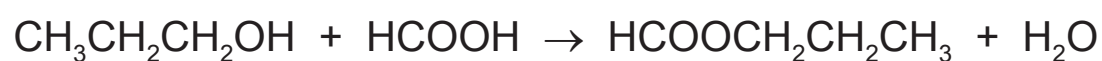
[2]

- (iii) Propanol is oxidised to a carboxylic acid by acidified potassium manganate(VII).

Deduce the name of this acid.

..... [1]

- (d) Propanol reacts with methanoic acid to form the ester propyl methanoate.



4.0g of methanoic acid was reacted with 6.0g of propanol.

- (i) Calculate the M_r of methanoic acid = [1]

- (ii) Calculate the M_r of propanol = [1]

- (iii) Determine which one is the limiting reagent. Show your reasoning.

.....

.....

..... [2]

- (iv) Calculate the maximum yield in grams of propyl methanoate, $M_r = 88$.

..... [1]

[Total: 17]

5 Iron is extracted from its ore, hematite, in a blast furnace.

Substances added to the furnace are:

- iron ore, hematite, containing impurities such as silica, SiO_2
- air
- coke, C
- limestone, CaCO_3

Substances formed in the blast furnace are:

- molten iron
- molten slag
- waste gases such as carbon dioxide

(a) State the **two** functions of the coke used in the blast furnace.

.....
 [2]

(b) Write an equation for the conversion of hematite, Fe_2O_3 , to iron.

..... [2]

(c) Explain how the silica impurity is removed and separated from the molten iron.

.....

 [3]

(d) The molten iron from the furnace is impure.
 It contains impurities which include the element carbon.

Explain how the carbon is removed. Include an equation in your answer.

.....

 [3]

[Total: 10]

- 6 The table below shows the elements in the third period of the Periodic Table, the number of electrons in their outer energy level, their oxidation state in their common compounds and their melting points.

element	Na	Mg	Al	Si	P	S	Cl	Ar
number of outer electrons	1	2	3	4	5	6	7	8
oxidation state	+1	+2	+3	+4/−4	−3	−2	−1	0
melting point/°C	98	650	660	1414	317	115	−101	−189

- (a) Describe and explain the variation in oxidation state across the period.

.....

 [3]

- (b) The first three elements, Na, Mg and Al, are metals.

Describe the structure of a typical metal.

.....

 [3]

- (c) Explain why Na, Mg and Al are good conductors of electricity.

..... [1]

- (d) Which element exists as diatomic molecules of the type X_2 ?

..... [1]

- (e) Silicon has a similar structure to diamond.

Explain why silicon has the highest melting point in the period.

.....
 [2]

- (f) Sodium chloride is a crystalline solid with a high melting point. It dissolves in water to give a neutral solution. Phosphorus trichloride is a liquid at room temperature. It reacts with water to form an acidic solution.

Suggest an explanation for these differences in properties.

.....
.....
..... [2]

- (g) Describe how you could show that magnesium oxide is a basic oxide and not an amphoteric oxide.

.....
..... [2]

- (h) Draw a dot-and-cross diagram showing the bonding in magnesium oxide. Show outer electrons only.

[3]

[Total: 17]

DATA SHEET
The Periodic Table of the Elements

		Group																	
		I	II	III	IV	V	VI	VII	O										
		<div style="display: flex; justify-content: space-around; align-items: center;"> <div style="text-align: center;"> ¹ H Hydrogen 1 </div> <div style="text-align: center;"> ² He Helium 2 </div> </div>																	
3	4	7 Li Lithium	9 Be Beryllium											19 F Fluorine	10 Ne Neon				
11	12	23 Na Sodium	24 Mg Magnesium	11 B Boron	6 C Carbon	7 N Nitrogen	8 O Oxygen	9 F Fluorine	10 Ne Neon	13 Al Aluminium	14 Si Silicon	15 P Phosphorus	16 S Sulfur	17 Cl Chlorine	18 Ar Argon				
19	20	39 K Potassium	40 Ca Calcium	21 Sc Scandium	22 Ti Titanium	23 V Vanadium	24 Cr Chromium	25 Mn Manganese	26 Fe Iron	27 Co Cobalt	28 Ni Nickel	29 Cu Copper	30 Zn Zinc	31 Ga Gallium	32 Ge Germanium	33 As Arsenic	34 Se Selenium	35 Br Bromine	36 Kr Krypton
37	38	85 Rb Rubidium	88 Sr Strontium	39 Y Yttrium	40 Zr Zirconium	41 Nb Niobium	42 Mo Molybdenum	43 Tc Technetium	44 Ru Ruthenium	45 Rh Rhodium	46 Pd Palladium	47 Ag Silver	48 Cd Cadmium	49 In Indium	50 Sn Tin	51 Sb Antimony	52 Te Tellurium	53 I Iodine	54 Xe Xenon
55	56	133 Cs Caesium	137 Ba Barium	57 La Lanthanum	72 Hf Hafnium	73 Ta Tantalum	74 W Tungsten	75 Re Rhenium	76 Os Osmium	77 Ir Iridium	78 Pt Platinum	79 Au Gold	80 Hg Mercury	81 Tl Thallium	82 Pb Lead	83 Bi Bismuth	84 Po Polonium	85 At Astatine	86 Rn Radon
87	88	226 Fr Francium	226 Ra Radium	89 Ac Actinium											85 At Astatine	86 Rn Radon			
		<p>*58-71 Lanthanoid series</p> <p>†90-103 Actinoid series</p>										165 Ho Holmium	102 No Nobelium						
		<div style="display: flex; justify-content: space-between;"> <div style="text-align: center;"> ^a X ^b </div> <div style="text-align: center;"> a = relative atomic mass X = atomic symbol b = proton (atomic) number </div> </div>										167 Er Erbium	101 Md Mendelevium						
		<div style="display: flex; justify-content: space-between;"> <div style="text-align: center;"> ^a X ^b </div> <div style="text-align: center;"> a = relative atomic mass X = atomic symbol b = proton (atomic) number </div> </div>										169 Tm Thulium	100 Fm Fermium						
		<div style="display: flex; justify-content: space-between;"> <div style="text-align: center;"> ^a X ^b </div> <div style="text-align: center;"> a = relative atomic mass X = atomic symbol b = proton (atomic) number </div> </div>										173 Yb Ytterbium	103 Lr Lawrencium						
		<div style="display: flex; justify-content: space-between;"> <div style="text-align: center;"> ^a X ^b </div> <div style="text-align: center;"> a = relative atomic mass X = atomic symbol b = proton (atomic) number </div> </div>										175 Lu Lutetium	103 Lr Lawrencium						

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).