

# UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS International General Certificate of Secondary Education

PHYSICAL SCIENCE 0652/11

Paper 1 Multiple Choice October/November 2012

45 minutes

Additional Materials: Multiple Choice Answer Sheet

Soft clean eraser

Soft pencil (type B or HB is recommended)

#### **READ THESE INSTRUCTIONS FIRST**

Write in soft pencil.

Do not use staples, paper clips, highlighters, glue or correction fluid.

Write your name, Centre number and candidate number on the Answer Sheet in the spaces provided unless this has been done for you.

There are **forty** questions on this paper. Answer **all** questions. For each question there are four possible answers **A**, **B**, **C** and **D**.

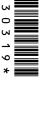
Choose the one you consider correct and record your choice in soft pencil on the separate Answer Sheet.

#### Read the instructions on the Answer Sheet very carefully.

Each correct answer will score one mark. A mark will not be deducted for a wrong answer.

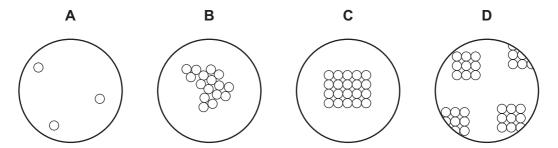
Any rough working should be done in this booklet.

A copy of the Periodic Table is printed on page 16.



**International Examinations** 

1 Which diagram shows the arrangement of particles in a liquid?



- 2 Which method can be used to obtain crystals from aqueous copper(II) sulfate?
  - **A** diluting
  - **B** dissolving
  - C evaporating
  - stirring
- 3 Statements 1, 2 and 3 are about diamond and graphite.
  - They are different solid forms of the same element.
  - 2 They each conduct electricity.
  - 3 They have atoms that form four equally strong bonds.

Which statements are correct?

- 1 only
  - **B** 3 only
- **C** 1 and 3
- **D** 2 and 3
- What is different for isotopes of the same element?
  - A number of electrons
  - В number of full shells
  - C number of nucleons
  - **D** number of protons
- Which compound has the largest relative molecular mass,  $M_r$ ?
  - $CO_2$
- B NO<sub>2</sub>
- C SiO<sub>2</sub>
- $SO_2$

**6** The equation below shows the reaction that occurs when hematite is heated with carbon.

process X

hematite + carbon 
$$\longrightarrow$$
 iron + carbon dioxide

 $2Fe_2O_3 + 3C$ 
 $4Fe + 3CO_2$ 

What is the chemical name of hematite and what is process X?

	chemical name	process X
Α	iron(II) oxide	oxidation
В	iron(II) oxide	reduction
С	iron(III) oxide	oxidation
D	iron(III) oxide	reduction

7 Magnesium reacts with acids to produce hydrogen gas.

Under which set of conditions is hydrogen produced most slowly?

	magnesium	acid	temperature/°C
Α	ribbon	concentrated	40
В	ribbon	dilute	20
С	powder	concentrated	40
D	powder	dilute	20

8 The chart shows the colour of Universal Indicator at different pH values.

colour red		(	oran	ge	green				blue				violet	
рН	1	2	3	4	5	6	7	8	9	10	11	12	13	14

Lemon juice contains citric acid which is only slightly acidic.

What colour does lemon juice give with Universal Indicator?

- A blue
- **B** green
- **C** orange
- **D** red

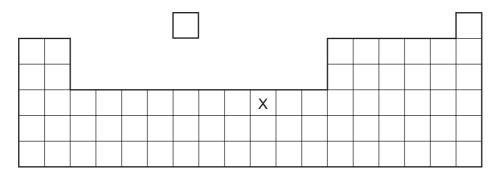
Aqueous ammonia is added to a solution of a metal sulfate. 9

A green precipitate forms that is insoluble in excess of the aqueous ammonia.

Which metal ion is present?

- A Cu<sup>2+</sup>
- **B**  $Fe^{2+}$  **C**  $Fe^{3+}$
- **D** Zn<sup>2+</sup>

**10** The position of an element, X, in the Periodic Table is shown.



Which correctly describes X?

	density (g/dm³)	melting point (°C)
Α	0.97	98
В	1.96	119
С	3.12	<b>–</b> 7
D	8.90	1455

11 Metal M is formed when its oxide is heated with carbon.

Which deductions from this information are correct?

- 1 M is similar in reactivity to iron.
- 2 M is more reactive than potassium.
- 3 The oxide of M is acidic.
- A 1 only
- **B** 1 and 3 only **C** 2 only **D** 2 and 3 only

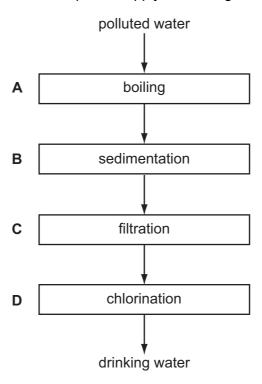
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**12** Copper, iron and zinc are all used to make things.

Which of these three metals are also used in the form of alloys?

	copper	iron	zinc
Α	<b>√</b>	<b>√</b>	<b>√</b>
В	✓	✓	X
С	X	✓	✓
D	X	X	✓

13 Which stage is **not** used to obtain the public supply of drinking water from polluted water?

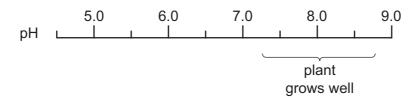


**14** In some reactions, carbon dioxide and water are both formed.

For which examples below is this statement correct?

- 1 burning of coal
- 2 reaction between an acid and a carbonate
- 3 respiration
- **A** 1 and 2 only **B** 1, 2 and 3 **C** 1 and 3 only **D** 2 and 3 only

15 The diagram shows the pH range of soil in which a certain plant grows well.



The plant is to be grown in a field with a soil pH of 6.

What can be added to the soil to make the pH suitable?

- A lime
- **B** litmus
- C nitric acid
- **D** sodium chloride

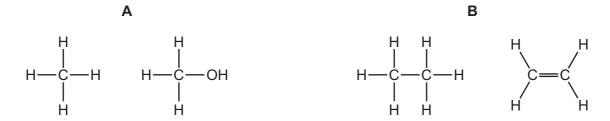
### **16** Which structure is **not** correct?

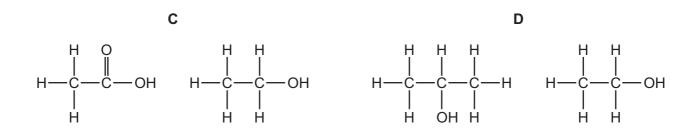
- 17 Three carbon-containing fuels are listed below.
  - 1 coal
  - 2 natural gas
  - 3 petroleum

Which of these fuels are classified as 'fossil fuels' and which are fractionally distilled?

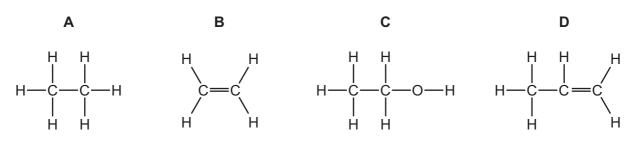
	fossil fuels	fractionally distilled
Α	1, 2 and 3	1 and 3 only
В	1, 2 and 3	3 only
С	1 and 3 only	1 and 3 only
D	1 and 3 only	3 only

18 Which two substances are in the same homologous series?

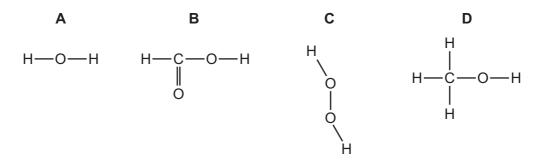




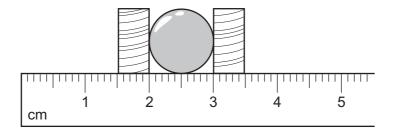
19 Which compound is the monomer used to make poly(ethene)?



20 Which molecular structure shows an alcohol?



21 A student uses two blocks and a ruler to find the radius of a ball.

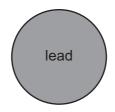


What is the radius of the ball?

- **A** 0.5 cm
- **B** 1.0 cm
- **C** 2.0 cm
- **D** 3.0 cm

**22** Three balls made of different materials are dropped from a bench.



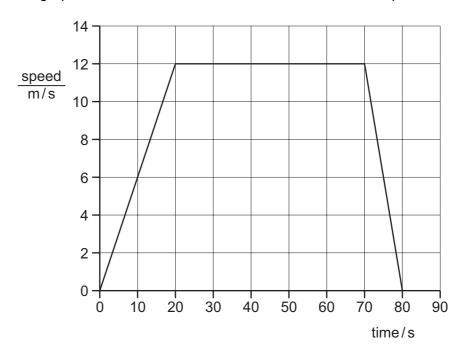




Which balls fall with the same acceleration?

- A aluminium and lead only
- B aluminium and wood only
- **C** lead and wood only
- **D** aluminium, lead and wood

23 The speed/time graph shown is for a bus as it travels from one bus stop to the next.



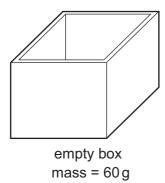
How far apart are the two bus stops?

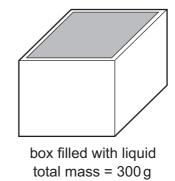
- **A** 120 m
- **B** 600 m
- **C** 780 m
- **D** 960 m

24 What is the unit of weight?

- A joule
- **B** kilogram
- C newton
- **D** watt

25 The diagrams show a rectangular box empty and filled with liquid.





The box has a mass of  $60\,g$  when empty. When filled with a liquid, the total mass of the box and the liquid is  $300\,g$ . The density of the liquid is  $1.2\,g/cm^3$ .

What is the volume of the liquid in the box?

- $\mathbf{A}$  50 cm<sup>3</sup>
- **B** 200 cm<sup>3</sup>
- **C** 250 cm<sup>3</sup>
- **D** 300 cm<sup>3</sup>

26 Which property of an object cannot be changed by a force?

- A its mass
- B its motion
- C its shape
- **D** its size

27 Which energy source stores gravitational energy?

- A coal
- **B** geothermal
- C hydroelectric
- **D** nuclear

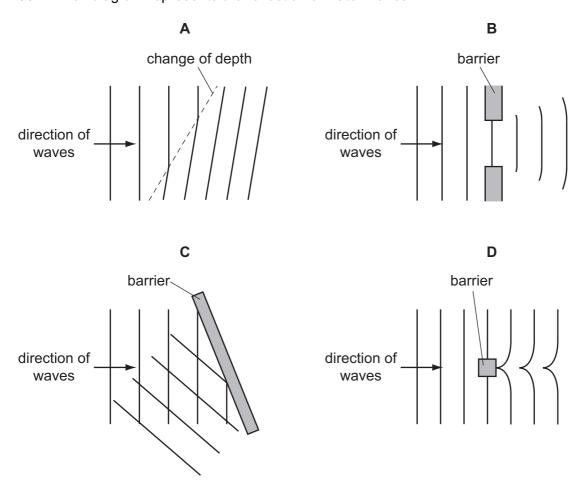
28 A car starts from rest and climbs a hill.

At the top of the hill, the car has gained 200 000 J of gravitational energy and 25 000 J of energy of motion. The thermal energy of the car and the surroundings has increased by 100 000 J.

How much chemical energy is used by the car?

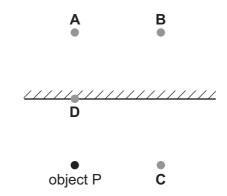
- **A** 125 000 J
- **B** 225 000 J
- **C** 300 000 J
- **D** 325 000 J

- 29 Which process involves convection?
  - A bread toasting under a grill
  - **B** heat energy passing through a copper bar
  - C heat from the Sun warming a road surface
  - **D** hot air rising to the top of a cool room
- 30 Which diagram represents the reflection of water waves?

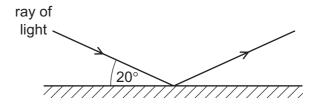


**31** A small object P is placed in front of a plane mirror as shown.

Where is the image of P formed?



**32** A ray of light strikes a plane mirror and reflects. The angle between the ray of light and the mirror is 20°.



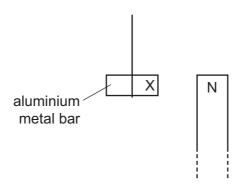
What is the size of the angle of reflection?

- **A** 20°
- **B** 70°
- **C** 140°
- **D** 160°

33 What is the approximate range of frequencies that can be heard by the human ear?

- **A** 1 Hz to 1000 Hz
- **B** 1 kHz to 1000 kHz
- **C** 20 Hz to 20 000 Hz
- **D** 20 kHz to 20 000 kHz

34 An aluminium bar is suspended near the north pole of a magnet.



What happens to the aluminium bar?

- **A** A north pole forms at X and the bar is attracted.
- **B** A north pole forms at X and the bar is repelled.
- **C** A south pole forms at X and the bar is attracted.
- **D** No pole forms at X and the bar is not affected.

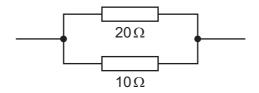
**35** An electric circuit contains a battery connected to a resistor.



Which values of electromotive force (e.m.f.) and resistance will produce the largest current?

	e.m.f./V	resistance/ $\Omega$
Α	3	5
В	3	10
С	12	40
D	12	80

**36** A  $20\Omega$  resistor and a  $10\Omega$  resistor are connected in parallel.



What is their combined resistance?

- **A** less than  $10\Omega$
- **B** 10Ω
- $\mathbf{C}$  20 $\Omega$
- **D** more than  $20\Omega$
- 37 The live, neutral and earth wires inside a mains lead are each covered by plastic insulation.

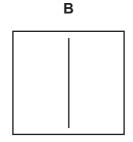
What is one purpose of the plastic?

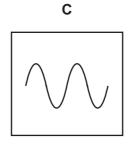
- **A** It increases the resistance of the wires.
- **B** It makes the wires stronger.
- **C** It stops current passing between the wires.
- **D** It stops heat escaping from the wires.

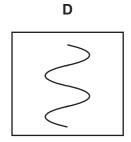
38 The diagrams show patterns which you might see on the screen of a cathode-ray oscilloscope.

Which pattern would appear if an alternating potential difference is applied to the Y-plates, with the time-base switched off?

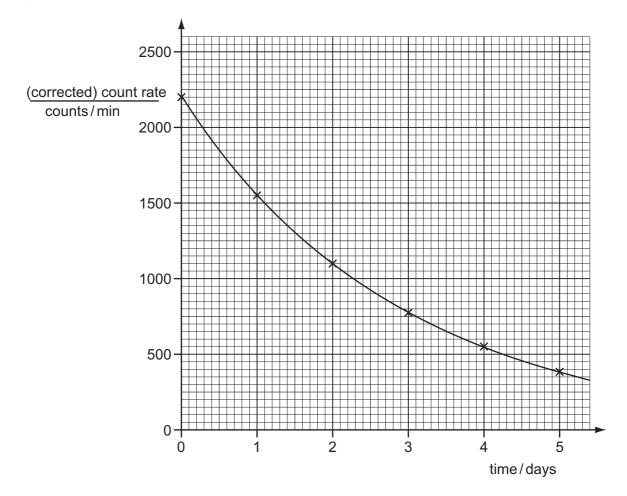
A \_\_\_\_\_







**39** The graph shows the decay curve for one particular radioactive isotope.



What is the half-life of this nuclide?

- **A** 1.0 day
- **B** 1.5 days
- **C** 2.0 days
- **D** 2.5 days

**40** A radium nuclide is represented by  ${}^{226}_{88}$ Ra.

How many nucleons are there in this nuclide?

- **A** 88
- **B** 138
- **C** 226
- **D** 314

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DATA SHEET
The Periodic Table of the Elements

	0	Heilum	20 <b>Ne</b> on	40 <b>Ar</b> Argon	84 X	36	£ <b>&gt;</b>	Xenon 54		Radon 86		175 <b>Lu</b> Lutetium 71	<b>Lr</b> Lawrencium
	II/		19 Fluorine	35.5 <b>C1</b> Chlorine	80 <b>Br</b>		127	_		Astatine 85		173 <b>Yb</b> Ytterbium 70	
			16 Oxygen	32 <b>S</b> Sulfur 16	79 <b>Se</b> Selenium	$\dashv$	128 <b>-</b>	E		Po Polonium 84		169 <b>Tm</b> Thulium	Mendelevium
	>		Nitrogen 8	31 Phosphorus	75 <b>As</b> Arsenic		122 <b>C.</b>		209	Bismuth 83		167 <b>Er</b> Erbium 68	Fm
	2		12 Carbon	28 <b>Si</b> Silicon			119		207			165 <b>Ho</b> Holmium 67	Einsteinium
	=		11 Boron 6	27 <b>A1</b> Auminium 13	70 <b>Ga</b> Gallium		115		204			Dy Dysprosium 66	
					65 <b>Zn</b> Zinc		112		201	Hg Mercury 80		159 <b>Tb</b> Terbium 65	<b>BK</b> Berkelium
					Copper			Silver 47		Au Good		Gd Gadolinium 64	Curium
dn					Signal Si	28	106	Palladium 46	195	Pt Platinum 78		152 <b>Eu</b> Europium 63	<b>Am</b> Americium
Group					59 Cobatt	27	103 <b>7</b>	Rhodium 45	192	<b>Ir</b> Iridium 77		Sm Samarium 62	<b>Pu</b> Plutonium
		T Hydrogen			56 <b>F.e.</b>	26	101	Ruthenium 44	190	Osmium Osmium 76		Pm Promethium 61	Neptunium
					55 Mn Manganese	25	Ę	E	186	<b>Re</b> Rhenium 75		Neodymium 60	238 <b>U</b>
					52 <b>Cr</b> Chromium	24	96 <b>2</b>	Ę	184	Tungsten 74		Pr Praseodymium 59	Pa Protactinium
					51 <b>V</b>	23	93	Niobium 41	181	<b>Ta</b> Tantalum 73		140 <b>Ce</b> Cerium	232 <b>Th</b>
					48 <b>T</b>	22	91	Zirconium 40	178	72			nic mass bol
					Scandium	21	% <b>&gt;</b>	Yttrium 39	139	Lanthanum 57 *	227 <b>Ac</b> Actinium 89	series eries	<ul> <li>a = relative atomic mass</li> <li>X = atomic symbol</li> <li>b = protein (atomic) number</li> </ul>
	=		Be Beryllium	24 Mg Magnesium 12	Calcium	20	88 0	Strontium 38	137	<b>Ba</b> Barium 56	226 <b>Ra</b> Radium 88	*58-71 Lanthanoid series	« <b>×</b>
	_		7 <b>Li</b> Lithium	23 <b>Na</b> Sodium	39 <b>K</b> Potassium	19	85	Rubidium 37	133	Caesium 55	<b>Fr</b> Francium 87	*58-71 L	Key

The volume of one mole of any gas is 24 dm<sup>3</sup> at room temperature and pressure (r.t.p.).

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