Cambridge Assessment International Education

Cambridge IGCSE™

CO-ORDINATED SCIENCES **0654/63**

Paper 6 Alternative to Practical **October/November 2020**

MARK SCHEME

Maximum Mark: 60

Published

This mark scheme is published as an aid to teachers and candidates, to indicate the requirements of the examination. It shows the basis on which Examiners were instructed to award marks. It does not indicate the details of the discussions that took place at an Examiners' meeting before marking began, which would have considered the acceptability of alternative answers.

Mark schemes should be read in conjunction with the question paper and the Principal Examiner Report for Teachers.

Cambridge International will not enter into discussions about these mark schemes.

Cambridge International is publishing the mark schemes for the October/November 2020 series for most Cambridge IGCSE™, Cambridge International A and AS Level and Cambridge Pre-U components, and some Cambridge O Level components.

[Turn over

PUBLISHED

Generic Marking Principles

These general marking principles must be applied by all examiners when marking candidate answers. They should be applied alongside the specific content of the mark scheme or generic level descriptors for a question. Each question paper and mark scheme will also comply with these marking principles.

GENERIC MARKING PRINCIPLE 1:

Marks must be awarded in line with:

- the specific content of the mark scheme or the generic level descriptors for the question
- the specific skills defined in the mark scheme or in the generic level descriptors for the question
- the standard of response required by a candidate as exemplified by the standardisation scripts.

GENERIC MARKING PRINCIPLE 2:

Marks awarded are always **whole marks** (not half marks, or other fractions).

GENERIC MARKING PRINCIPLE 3:

Marks must be awarded **positively**:

- marks are awarded for correct/valid answers, as defined in the mark scheme. However, credit is given for valid answers which go beyond the scope of the syllabus and mark scheme, referring to your Team Leader as appropriate
- marks are awarded when candidates clearly demonstrate what they know and can do
- marks are not deducted for errors
- marks are not deducted for omissions
- answers should only be judged on the quality of spelling, punctuation and grammar when these features are specifically assessed by the question as indicated by the mark scheme. The meaning, however, should be unambiguous.

GENERIC MARKING PRINCIPLE 4:

Rules must be applied consistently, e.g. in situations where candidates have not followed instructions or in the application of generic level descriptors.

GENERIC MARKING PRINCIPLE 5: Marks should be awarded using the full range of marks defined in the mark scheme for the question (however; the use of the full mark range may be limited according to the quality of the candidate responses seen).
GENERIC MARKING PRINCIPLE 6: Marks awarded are based solely on the requirements as defined in the mark scheme. Marks should not be awarded with grade thresholds or grade descriptors in mind.

Science-Specific Marking Principles

1	Examiners should consider the context and scientific use of any keywords when awarding marks. Although keywords may be present, marks should not be awarded if the keywords are used incorrectly.
2	The examiner should not choose between contradictory statements given in the same question part, and credit should not be awarded for any correct statement that is contradicted within the same question part. Wrong science that is irrelevant to the question should be ignored.
3	Although spellings do not have to be correct, spellings of syllabus terms must allow for clear and unambiguous separation from other syllabus terms with which they may be confused (e.g. ethane / ethene, glucagon / glycogen, refraction / reflection).
4	The error carried forward (ecf) principle should be applied, where appropriate. If an incorrect answer is subsequently used in a scientifically correct way, the candidate should be awarded these subsequent marking points. Further guidance will be included in the mark scheme where necessary and any exceptions to this general principle will be noted.
5	‘List rule’ guidance For questions that require ***n*** responses (e.g. State **two** reasons …): • The response should be read as continuous prose, even when numbered answer spaces are provided. • Any response marked *ignore* in the mark scheme should not count towards ***n***. • Incorrect responses should not be awarded credit but will still count towards ***n***. • Read the entire response to check for any responses that contradict those that would otherwise be credited. Credit should **not** be awarded for any responses that are contradicted within the rest of the response. Where two responses contradict one another, this should be treated as a single incorrect response. • Non-contradictory responses after the first ***n*** responses may be ignored even if they include incorrect science.

6 Calculation specific guidance

Correct answers to calculations should be given full credit even if there is no working or incorrect working, **unless** the question states 'show your working'.

For questions in which the number of significant figures required is not stated, credit should be awarded for correct answers when rounded by the examiner to the number of significant figures given in the mark scheme. This may not apply to measured values.

For answers given in standard form (e.g. $a \times 10^n$) in which the convention of restricting the value of the coefficient (a) to a value between 1 and 10 is not followed, credit may still be awarded if the answer can be converted to the answer given in the mark scheme.

Unless a separate mark is given for a unit, a missing or incorrect unit will normally mean that the final calculation mark is not awarded. Exceptions to this general principle will be noted in the mark scheme.

7 Guidance for chemical equations

Multiples / fractions of coefficients used in chemical equations are acceptable unless stated otherwise in the mark scheme.

State symbols given in an equation should be ignored unless asked for in the question or stated otherwise in the mark scheme.

Question	Answer	Marks
1(a)(i)	13 ;	**1**
1(b)(i)	Axes right way round and correctly labelled including units; scales linear and plotted points cover at least half the grid; all plots correct ± half small square;	**3**
1(b)(ii)	best fit curve judgement ;	**1**
1(c)	correct reading from graph; graph marked ;	**2**
1(d)(i)	same hydrogen peroxide / hydrogen peroxide not changed ; difficulty cutting slices; difficulty seeing when at top ; temperature ; no repeats / anomalies not identified	**1**
1(d)(ii)	depth / volume of hydrogen peroxide; concentration of hydrogen peroxide; temperature; same size beaker; same potato / species; diameter of slice;	**2**
1(e)	glowing splint; relights;	**2**

Question	Answer	Marks
2(a)(i)	blue-black; (pale) blue ;	**2**
2(a)(ii)	starch; protein ;	**2**
2(b)(i)	reducing sugar ;	**1**
2(b)(ii)	yellow / green; orange / red ;	**2**
2(c)	Benedict's ;	**1**

Question	Answer	Marks
3(a)	to dissolve (the soluble substance)	**1**
3(b)(i)	(carbon dioxide bubbled into) limewater goes milky ;	**1**
3(b)(ii)	zinc ;	**1**
3(c)	chloride / bromide / iodide ;	**1**
3(d)	yes gives carbon dioxide with acid ; yes gives a white precipitate with barium nitrate '	**2**

Question	Answer	Marks
4(a)(i)	the readings do not have a consistent number of decimal places / different number of sig figs ;	**1**
4(a)(ii)	1.02 ;	**1**
4(b)	the test-tube might break / and crack the glass ;	**1**
4(c)(i)	22.67 **AND** 22.56 ;	**1**
4(c)(ii)	decomposition is not finished because the mass decreases ;	**1**
4(c)(iii)	0.38 ;	**1**
4(c)(iv)	0.64 ;	**1**
4(d)	blue flame hotter / yellow flame decomposition will be slow / some soot will be deposited on the test tube / mass of carbon dioxide and water measured will be less than it should be / test-tube goes black ;	**1**

Question	Answer	Marks
5	**One mark for each section and any two other marking points** **apparatus** power supply and electrodes dipping in (dilute) sulfuric acid; use of carbon electrodes ; suitable apparatus for collecting known volume of gas at electrode ; stop-watch ; **method and measurements** measure volume / number bubbles ; fixed time ; use (at least five) different concentrations of dilute sulfuric acid ; use 5 concentrations of acid ; repeat every one ; **controlling variables** control the temperature ; control the current ; control size / mass electrode ; **processing and drawing conclusion** calculate average ; see if higher / lower concentration makes more / less gas / bubbles **;** plot graph of time / volume against concentration of sulfuric acid ; look for trend in graph ; calculate rate volume gas ÷ time ;	**6**

Question	Answer	Marks
6(a)	correct symbol ; correct parallel connection ;	**2**
6(b)	2.1 (V) 0.24 (A)	**2**
6(c)(i)	so that the power supply does not run down / wires get hot / resistance changes	**1**
6(c)(ii)	8.8 ; 2.1 ;	**2**
6(c)(iii)	Ω	**1**
6(d)	It is about 4 times / 9 times would be 18.9	**1**
6(e)	Circuit 2 / parallel AND the lamps are brighter	**1**

Question	Answer	Marks
7(a)(i)	81.0 ;	**1**
7(a)(ii)	to allow the thermometer to record the maximum temperature (of the hot water) ;	**1**
7(a)(iii)	s, °C, °C	**1**
7(a)(iv)	to ensure that all the water is at the same temperature ;	**1**
7(b)	11(.0) ;	**1**
7(c)	5(.0) ;	**1**
7(d)	the lid reduces the rate of fall in temperature or equivalent ;	**1**
7(e)	tongs / glove **AND** protect hands / skin **AND** from burns / hot water ; goggles **AND** protect eyes **AND** from hot water ;	**1**
7(f)	insulate / lag the beaker	**1**
7(g)	1 from: volume of water initial temperature of the hot water room temperature ;	**1**