

**OXFORD CAMBRIDGE AND RSA EXAMINATIONS
GCSE (9–1)**

J248/01

CHEMISTRY A (GATEWAY SCIENCE)

**Paper 1, C1–C3 and C7
(Foundation Tier)**

**THURSDAY 17 MAY 2018: Morning
TIME ALLOWED: 1 hour 45 minutes
plus your additional time allowance
MODIFIED ENLARGED 24pt**

First name		Last name	
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Centre number						Candidate number				
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YOU MUST HAVE:

a ruler (cm/mm)

the Data Sheet (for GCSE Chemistry A)

YOU MAY USE:

a scientific or graphical calculator

an HB pencil

Models for Question 21

READ INSTRUCTIONS OVERLEAF



INSTRUCTIONS

The data sheet will be found with this document.

Use black ink. You may use an HB pencil for graphs and diagrams.

Complete the boxes on the front page with your name, centre number and candidate number.

Answer ALL the questions.

Write your answer to each question in the space provided. If additional space is required, use the lined page(s) at the end of this booklet. The question number(s) must be clearly shown.

INFORMATION

The total mark for this paper is 90.

The marks for each question are shown in brackets [].

Quality of extended responses will be assessed in questions marked with an asterisk (*).

SECTION A

Answer ALL the questions.

You should spend a maximum of 30 minutes on this section.

1 Which of these pH values shows the pH of a strong acid? [1]

A 1

B 5

C 7

D 10

Your answer

2 Which of these general properties correctly describes a metal? [1]

A Ductile and good conductor of heat

B High density and forms negative ions

C Malleable and low density

D Shiny and brittle

Your answer

3 A number of scientists contributed to the development of the atomic model.

Which of these scientists discovered the electron? [1]

A Bohr

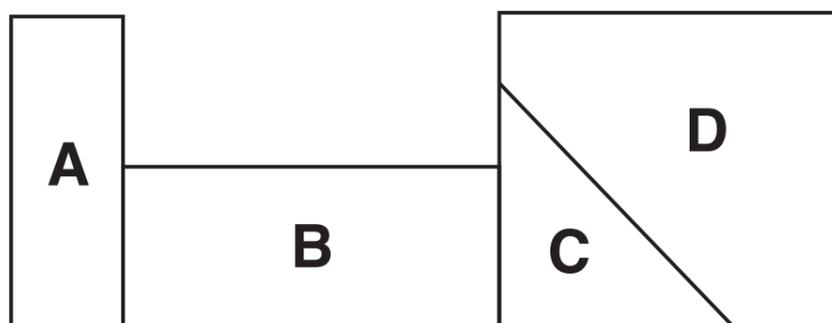
B Dalton

C Rutherford

D Thomson

Your answer

4 This is a section of the Periodic Table.



In which section of the Periodic Table would you find NON-METALS? [1]

Your answer

5 What is the typical size of the radius of an atom? [1]

A 10^{-2} m

B 10^{-5} m

C 10^{-10} m

D 10^{-20} m

Your answer

6 Lead nitrate contains lead ions, Pb^{2+} , and nitrate ions, NO_3^- .

What is the formula for lead nitrate? [1]

A PbNO_3

B $\text{Pb}(\text{NO}_3)_2$

C Pb_2NO_3

D $\text{Pb}_2(\text{NO}_3)_2$

Your answer

7 Look at the equation.



Which substance is **OXIDISED** in this reaction? [1]

A CH_4

B CO_2

C H_2O

D O_2

Your answer

8 Look at the equation.



Which substance is the **OXIDISING AGENT** in this reaction? [1]

A CH_4

B CO_2

C H_2O

D O_2

Your answer

9 Which statement about COVALENT bonding is true? [1]

A Electrons are transferred from one atom to another.

B Electrons are delocalised.

C Electrons are shared between atoms.

D Ions are formed.

Your answer

10 The electronic structure of an atom of an element is 2.8.8.2.

In which PERIOD of the Periodic Table is this element found? [1]

A 1

B 2

C 4

D 8

Your answer

11 The electronic structure of an atom of an element is 2.8.8.2.

In which GROUP of the Periodic Table is this element found? [1]

A 1

B 2

C 4

D 8

Your answer

12 What is the name of the gas made when zinc carbonate reacts with hydrochloric acid? [1]

A Carbon dioxide

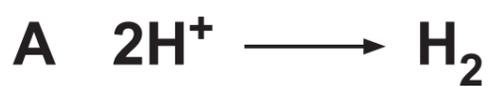
B Chlorine

C Hydrogen

D Oxygen

Your answer

13 Which equation represents NEUTRALISATION? [1]



Your answer

14 Which of these statements about nanoparticulate materials is correct? [1]

A Nanoparticles are much smaller than atoms.

B Nanoparticulate materials can be used as catalysts.

C Nanoparticulate materials have a very small surface area to volume ratio.

D There are no risks when using nanoparticulate materials.

Your answer

15 Ethanol is a liquid at room temperature. It has a low melting point and boiling point.

Why? [1]

A Ethanol is an ionic compound.

B The forces of attraction between ethanol molecules are strong.

C The forces of attraction between ethanol molecules are weak.

D There are no forces of attraction between ethanol molecules.

Your answer

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SECTION B

Answer ALL the questions.

16 A student investigates the energy changes during some chemical reactions.

She measures the temperature at the start and end of each reaction.

Look at her results.

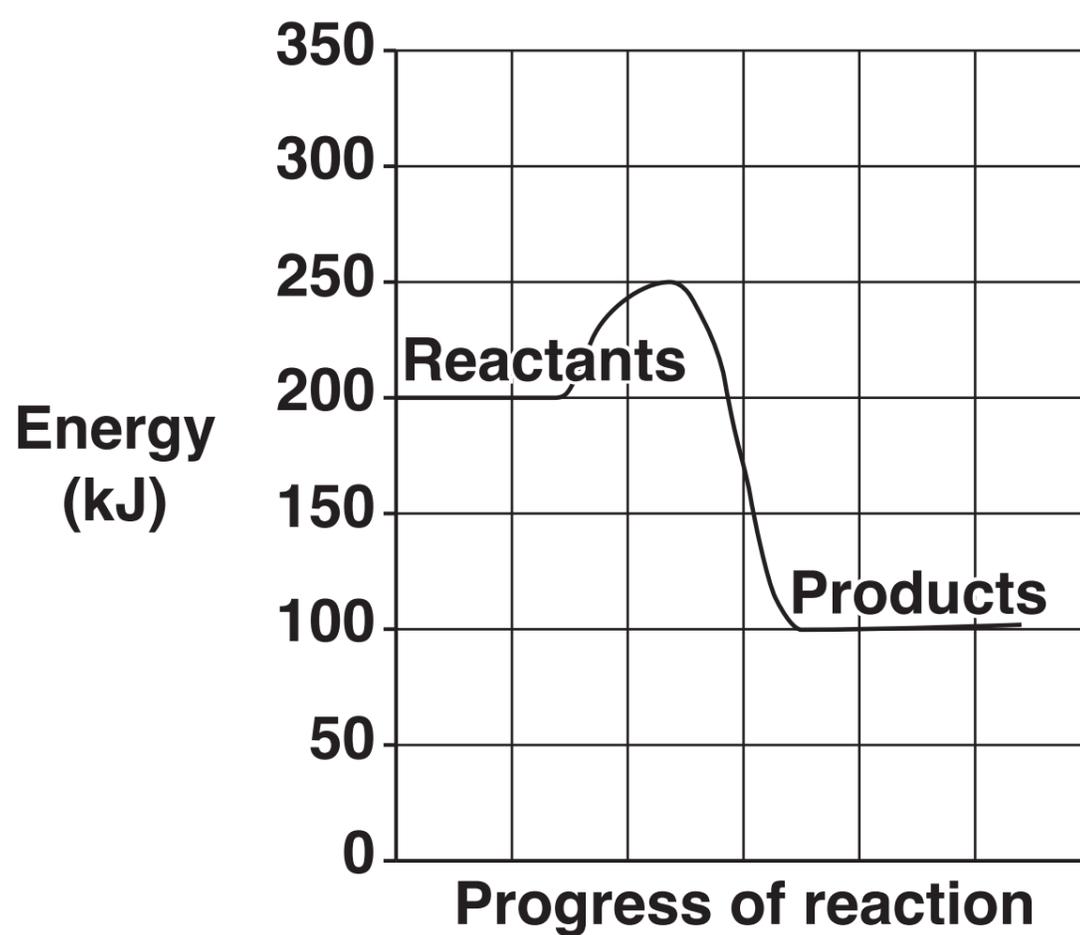
Reaction	Temperature at start (°C)	Temperature at end (°C)	Temperature change (°C)	Type of reaction
A	20	25		Exothermic
B	18	10		
C	21	35		
D	20	20		No reaction

(a) Complete the table. [3]

(b) Which reaction has the largest energy change?

Answer = _____ [1]

(c) Look at the reaction profile for reaction A.



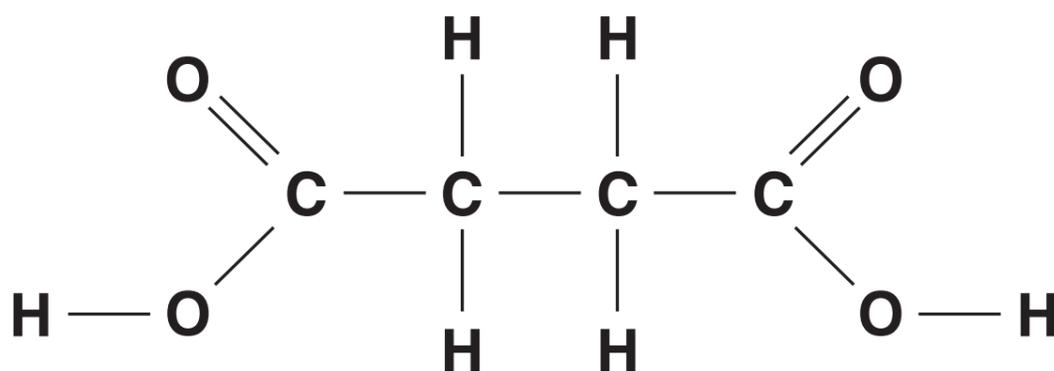
(i) Calculate the energy change in this reaction.

Answer = _____ kJ [1]

(ii) Calculate the activation energy.

Answer = _____ kJ [1]

17 Look at the diagram. It shows the displayed formula of succinic acid.



(a) Complete the table to show the number of atoms of each element in this displayed formula. [2]

Element	Number of atoms
C	
H	
O	

(b) What is the EMPIRICAL FORMULA of succinic acid?

_____ [1]

(c) Succinic acid has a melting point of 184°C and a boiling point of 235°C .

What is the state of succinic acid at 25°C ?

Explain your answer.

_____ [2]

18 A student is separating some mixtures. She wants to make pure water from a solution of salt water. She filters the mixture.

Her method does not work.

(a) Explain why her method does not work and describe the method she should use.

[2]

(b) The student wants to separate a mixture of two liquids.

The liquids are:

Liquid	Boiling point (°C)
Water	100
Ethanol	78

Which separation technique should she use?

Explain how the method works.

[2]

(c) The student separates two solid substances A and B. She wants to check that they are PURE.

(i) What is meant by a PURE SOLID?

_____ [1]

(ii) The student measures the melting points of four samples of solid A.

Look at her results.

Sample	Melting point (°C)
1	115
2	119
3	114–118
4	120–122

She knows that a pure sample of solid A has a melting point of 120 °C.

She concludes that sample 4 is the purest sample of solid A.

Do the results support her conclusion?

Explain your answer using evidence from the table.

_____ [3]

19 Two students heat some calcium carbonate, CaCO_3 .

Look at the equation for the reaction.



(a) What is the meaning of (s) in the equation?

[1]

(b) Look at their results.

Mass of calcium carbonate (g)	Mass of calcium oxide (g)	Mass of carbon dioxide (g)
1.00	0.56	0.44
2.00	1.12	0.88
3.00	1.68	1.32
4.00	2.24	

Complete the table. [1]

(c) Student A states:

'If I heat 20 g of calcium carbonate, I will make 8.8 g of calcium oxide and 11.2 g of carbon dioxide.'

Is student A correct?

Explain your answer.

[2]

(d) Student B investigates another reaction.

Look at the equations.



magnesium + oxygen \longrightarrow magnesium oxide

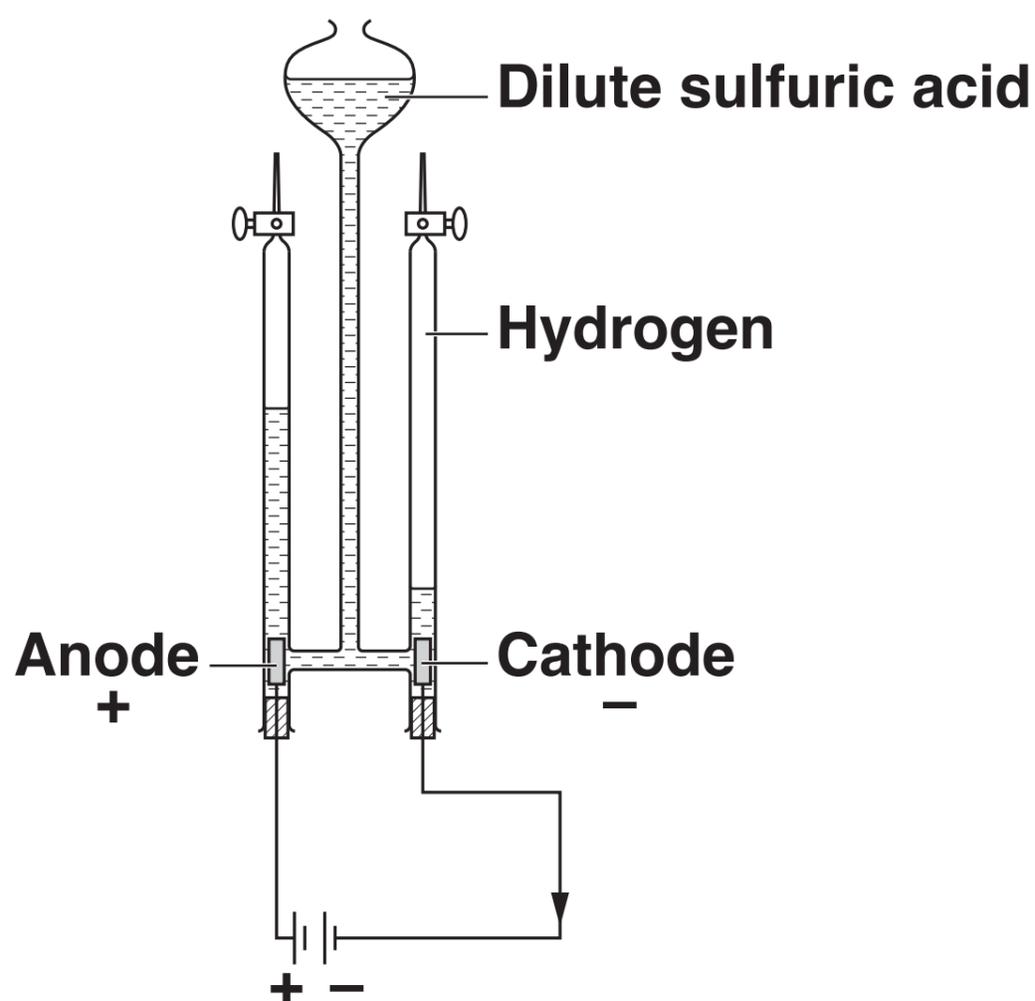
(i) Calculate the relative formula mass of magnesium oxide.

Answer = _____ [1]

(ii) Use the relative formula mass of magnesium oxide and the relative atomic masses of magnesium and oxygen to show if mass is conserved during this reaction.

_____ [2]

20 A student electrolyses dilute sulfuric acid.



Hydrogen gas is made at the cathode.

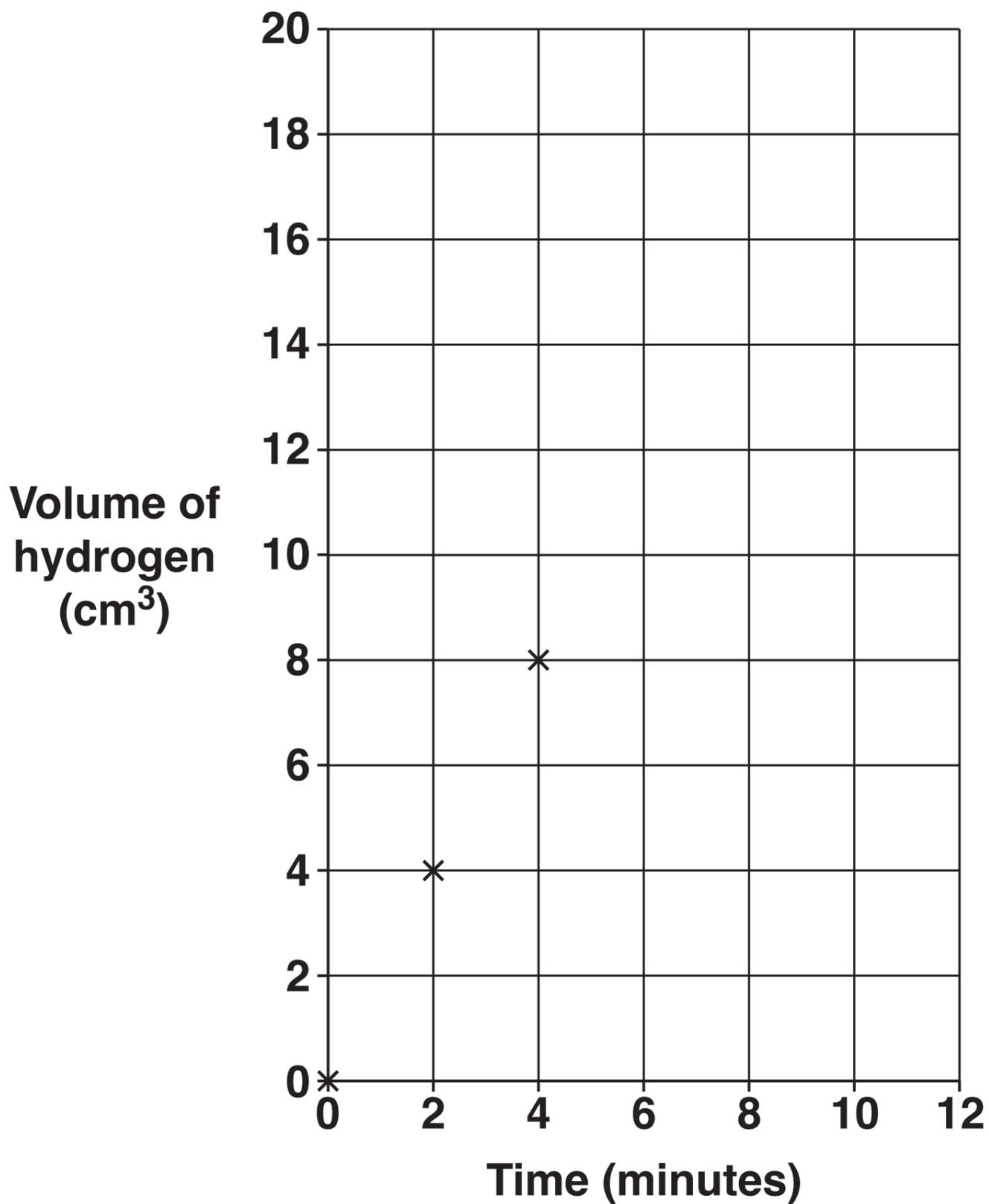
The student measures the volume of hydrogen made at the cathode every 2 minutes for 10 minutes.

Look at his results.

Time (minutes)	Volume of hydrogen (cm ³)
0	0.0
2	4.0
4	8.0
6	14.0
8	16.0
10	20.0

(a) Plot the results on the grid. The first 3 points have been done for you.

Draw a line of best fit. [2]



(b) One of the results is ANOMALOUS.

Circle the anomalous result on the graph. [1]

(c) Sulfuric acid contains these particles.



Which particles are attracted to the ANODE?

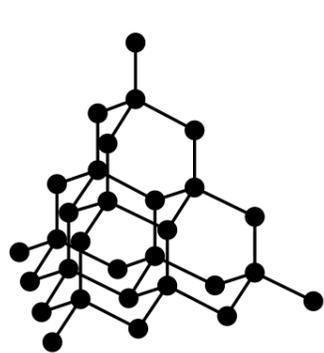
_____ [1]

(d) The student also investigates the electrolysis of some molten (liquid) salts.

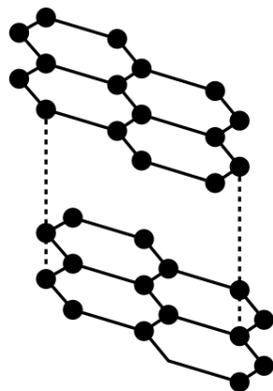
Complete the table. [2]

Molten salt	Formula	Product at cathode	Product at anode
Potassium chloride	KCl	Potassium	
Lead iodide	PbI_2		Iodine

21 The diagrams show the structures of diamond and graphite. You may also use models to help you.



Diamond



Graphite

One property of diamond is that it is very hard. One property of graphite is that it is slippery.

(a) Write about the other properties of diamond and graphite.

Diamond _____

Graphite _____

[4]

(b) Describe the type of bonding between the carbon atoms in diamond.

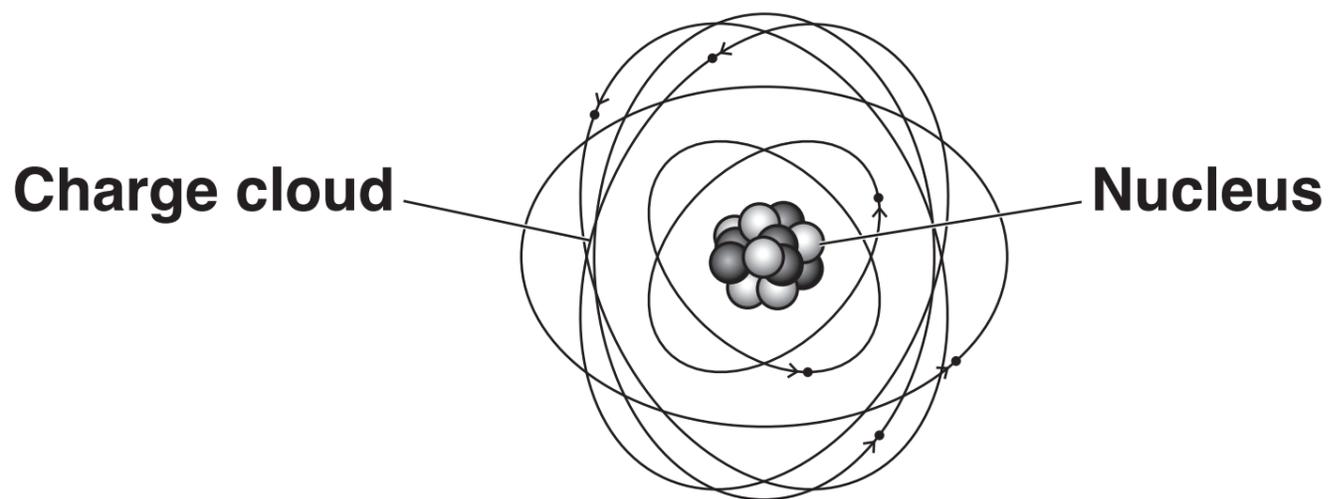
[1]

(c) Graphite is slippery.

Use the structure of graphite to explain why.

[2]

22 Look at the diagram of an atom.



(a) Which particles are present in the charge cloud?

_____ [1]

(b) Which TWO particles make up the nucleus?

and _____ [1]

(c) Most of the mass of an atom is in the nucleus.

Explain why.

_____ [2]

(d) Look at these two atoms of chlorine.



What is the relationship between these two atoms of chlorine?

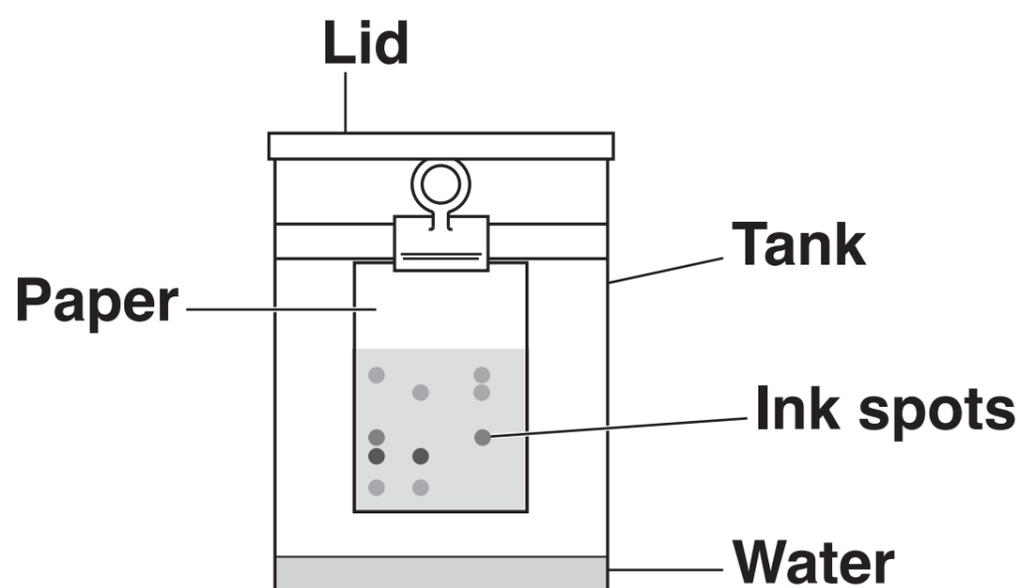
Explain your answer.

[2]

23 A forensic scientist is investigating the ink that has been used to forge the signature on a cheque.

She separates the colours in some inks using paper chromatography.

Look at the diagram of her apparatus.



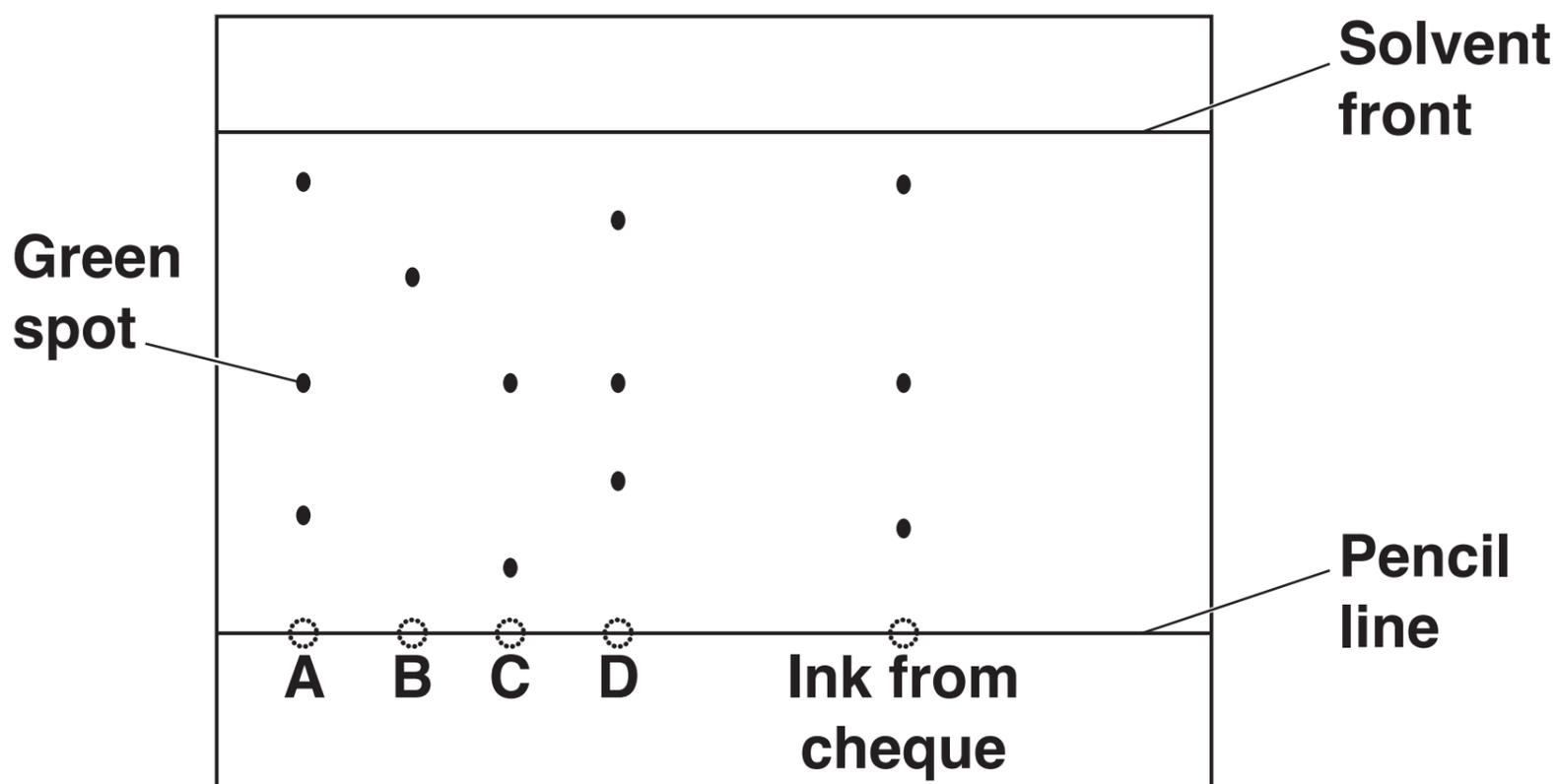
(a) What is the MOBILE PHASE in this experiment?

_____ [1]

(b) Explain how paper chromatography separates the colours in ink.

_____ [1]

(c) Look at the results of the scientist's experiment.



(i) Look at the green spot for ink A.

Calculate the R_f value for the green spot.

Answer = _____ [2]

(ii) Which ink was used to forge the signature on the cheque?

Explain how you can tell.

[2]

25 Magnesium is an element. It is solid at room temperature.

(a) (i) SOLID magnesium cannot be compressed.

Why?

_____ [1]

(ii) SOLID magnesium CANNOT flow, but LIQUID magnesium CAN flow.

Explain why.

_____ [3]

(iii) Magnesium GAS completely fills any container it is put in.

Explain why.

_____ [2]

(b) Magnesium reacts with water. Magnesium hydroxide, $\text{Mg}(\text{OH})_2$, and hydrogen, H_2 , are made.

Write a balanced symbol equation for this reaction.

_____ [2]

(c) Magnesium nitrate has the formula $\text{Mg}(\text{NO}_3)_2$.

Calculate the relative formula mass of magnesium nitrate.

Answer = _____ [1]

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(c) The student also investigates other reactions.

The table shows the salts he can make from different starting materials.

Complete the table. [3]

Acid used	Other starting material	Salt made
Sulfuric acid	Copper oxide	
	Zinc carbonate	Zinc nitrate
Hydrochloric acid		Magnesium chloride

(d) What TYPE of reaction happens when sulfuric acid reacts with copper oxide?

_____ [1]

END OF QUESTION PAPER

