



Mark Scheme (Results)

January 2017

International GCSE
Chemistry (4CH0) Paper 2C

Pearson Edexcel Certificate in
Chemistry (KCH0) Paper 2C

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General Marking Guidance

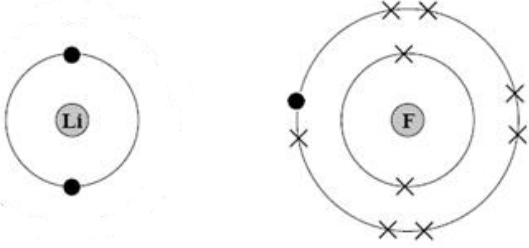
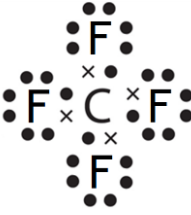
- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

Question number	Answer	Notes	Marks
1 (a)	air		1
(b)	hydrogen	ACCEPT H ₂ IGNORE H	
(c)	chlorine	ACCEPT Cl ₂ IGNORE Cl	1
(d)	chlorine	ACCEPT Cl ₂ IGNORE Cl	1
(e)	iron	ACCEPT Fe	1
		Total	5

Question number	Answer	Notes	Marks
2 (a) (i)	(both are) solids AND (both) form alkaline solutions (in water)	ALLOW (both are) <u>slightly</u> soluble	1
(ii)	(both are) gases AND (both) form acidic solutions (in water)		1
(b) (i)	the oxide is solid		1
(ii)	the oxide forms an acidic solution (in water)		1
(c)	M1 the lamp does not light up M2 (so this shows that) phosphorus/it does not conduct electricity	ACCEPT reverse arguments	2
		Total	6

Question number	Answer	Notes	Marks
3 (a)	<p>M1 calcium is the most reactive; titanium is the second most reactive; tin is the least reactive</p> <p>OR</p> <p>calcium most reactive AND tin least reactive</p> <p>M2 (because) titanium displaces tin</p> <p>M3 (and because) titanium does not displace calcium</p>	<p>ACCEPT Ca > Ti > Sn</p> <p>ACCEPT replaces</p> <p>ACCEPT replaces</p>	3
(b) (i)	$2\text{Al} + \text{Fe}_2\text{O}_3 \rightarrow \text{Al}_2\text{O}_3 + 2\text{Fe}$	ACCEPT multiples and halves	1
(ii)	<p>M1 aluminium</p> <p>M2 because it has gained oxygen</p> <p>OR</p> <p>because it has lost electrons</p> <p>OR</p> <p>because its oxidation number has increased</p> <p>OR</p> <p>because its oxidation number has changed from 0 to +3</p>	M2 DEP on M1	2
(iii)	<p>M1 (powders have) larger surface area</p> <p>M2 (therefore) faster reaction</p>	<p>ACCEPT reverse arguments</p> <p>IGNORE references to collisions between particles</p>	2
		Total	8

Question number	Answer	Notes	Marks
4 (a)	D (NH_4^+)		1
(b)	D (I^-)		1
(c)	A (carbonate)		1
(d)	C (blue, green, brown)		1
(e)	A (carbonate)		1
		Total	5

Question number	Answer	Notes	Marks
5 (a)		<p>1 mark for each correct ion</p> <p>IGNORE charges on ion</p> <p>ALLOW any combination of dots and crosses</p> <p>Diagram showing sharing of electrons scores 0</p>	2
(b)		<p>M1 all four bonding pairs correct, and only 8 electrons shown in outer shell of carbon</p> <p>M2 all non-bonding pairs correct</p> <p>IGNORE inner shells even if incorrect</p> <p>If rings are drawn then both bonding electrons must be in the overlapping area</p> <p>M2 DEP on M1</p>	2

Question number	Answer	Notes	Marks
(c)	<p>M1 (high melting point of LiF) strong forces (of attraction) between (oppositely charged) ions / strong forces (of attraction) between Li^+ and F^-</p> <p>M2 (good conductivity of LiF) ions are mobile</p> <p>M3 (low melting point of CF_4) weak forces (of attraction) between molecules / weak intermolecular forces (of attraction)</p> <p>M4 (poor conductivity of CF_4) molecules are not charged / molecules are neutral / no charged particles</p>	<p>ACCEPT strong (ionic) bonding / strong (ionic) bonds ACCEPT large amount of energy required to overcome the forces between the ions/break (ionic) bonds REJECT any reference to molecules or intermolecular forces REJECT reference to atoms</p> <p>ACCEPT ions can move REJECT any reference to electrons are mobile/delocalised electrons</p> <p>ACCEPT weak van der Waals forces / weak London forces / weak dispersion forces ACCEPT very little energy required to overcome the intermolecular forces. ALLOW weak intermolecular bonds REJECT any references to covalent bonds broken</p> <p>IGNORE no free electrons/all electrons used in bonding IGNORE there are no ions ALLOW no delocalised / no mobile electrons</p>	4
		Total	8

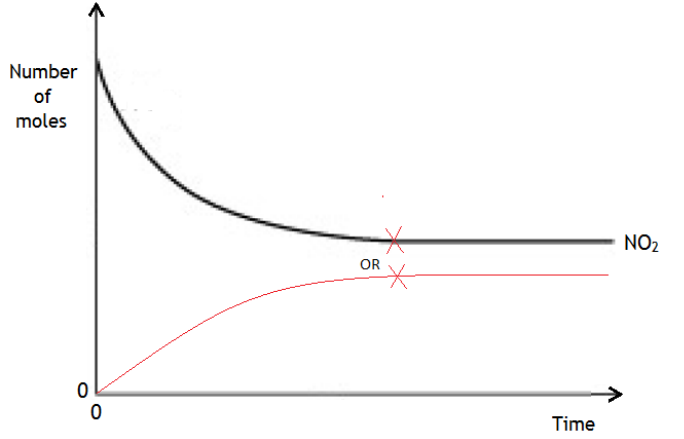
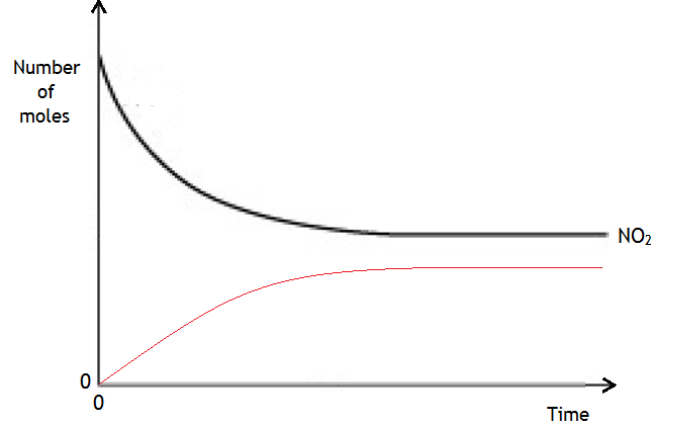
Question number	Answer	Notes	Marks
6 (a) (i)	(to provide the) zymase/enzyme (that acts as a catalyst)	ALLOW (to act as a) catalyst ALLOW to increase the rate of reaction IGNORE to lower the activation energy IGNORE to start the reaction REJECT to provide (activation) energy	1
(ii)	(turns) milky / cloudy / turbid (then clear) OR <u>white</u> precipitate / <u>white</u> suspension / <u>white</u> solid (forms then disappears)		1
(iii)	30 °C	ACCEPT any temperature, or range of temperatures, between 25 and 40 °C	1

Question number	Answer	Notes	Marks
(b)	<p>Route 1:</p> <p>M1 $\Sigma(\text{bonds broken}) =$ $348 + (5 \times 412) + 360 + 463 + (3 \times 496)$ OR 4719 (kJ/mol)</p> <p>M2 $\Sigma(\text{bonds made}) = (4 \times 743) + (6 \times 463)$ OR 5750 (kJ/mol)</p> <p>Route 2:</p> <p>M1 $\Sigma(\text{bonds broken}) =$ $348 + (5 \times 412) + 360 + (3 \times 496)$ OR 4256 (kJ/mol)</p> <p>M2 $\Sigma(\text{bonds made}) = (4 \times 743) + (5 \times 463)$ OR 5287 (kJ/mol)</p> <p>M3 $4719 - 5750 \text{ (kJ/mol) / M1 - M2}$</p> <p>M4 $- 1031 \text{ (kJ/mol)}$ OR correct evaluation of M3</p>	<p>IGNORE negative sign</p> <p>IGNORE negative sign</p> <p>Sign required ACCEPT answers given to three significant figures</p> <p>Correct answer with no working scores 4 + 1031 (kJ/mol) scores 3</p>	4

Question number	Answer	Notes	Marks
6 (c) (i)	<p>M1 32</p> <p>M2 $(32 \times 15.6) = 500$ (kJ)</p> <p>OR</p> <p>M1 $\times 15.6$ correctly evaluated</p>	<p>ACCEPT 499 / 499.2</p> <p>Correct answer with no working scores 2</p>	2
(ii)	<p>M1 & M2 Any two from:</p> <ul style="list-style-type: none"> • mass of water / volume of water / amount of water • distance of flame from the can • length of wick 	<p>IGNORE temperature of water (at start)</p> <p>ALLOW distance of burner from can</p>	2
(iii)	<p>Any two from:</p> <p>M1 heat (energy)/thermal energy is lost (to /surroundings)</p> <p>M2 incomplete combustion (of the fuel)</p> <p>M3 evaporation of water/fuel</p>	<p>IGNORE just energy lost</p> <p>IGNORE not all of the ethanol is burned</p>	2
		Total	13

Question number	Answer	Notes	Marks
7 (a) (i)	<p>M1 0.080 mol of HCl react with 0.040 mol of MgCO₃</p> <p>M2 0.050 > 0.040</p>	<p>ACCEPT any method involving correct ratios of moles, eg HCl to MgCO₃ is 2:1 0.08 to 0.05 is 2:1.25</p> <p>ACCEPT correct calculations involving masses</p>	2
(ii)	<p>M1 $n(\text{CO}_2) = \frac{1}{2} \times 0.08(0)$ OR $0.04(0)$</p> <p>M2 $\text{vol}(\text{CO}_2) = (0.04(0) \times 24\,000)$ $= 960 \text{ (cm}^3\text{)}$</p> <p>OR</p> <p>M1 $\times 24\,000$ correctly evaluated</p>	<p>Correct answer with no working scores 2</p> <p>1920 (cm³) scores 1 mark</p>	2
(b) (i)	<p>M1 $M_r(\text{MgCl}_2 \cdot 6\text{H}_2\text{O}) = 203$</p> <p>M2 % yield = $((5.5 \div (203 \times 0.050)) \times 100)$ $= 54$</p>	<p>M2 CSQ on M1</p> <p>ACCEPT any number of significant figures except one (eg reject 50)</p> <p>Calculator value is 54.1871921182</p> <p>REJECT answers > 100</p>	2
(ii)	<p>Any one of:</p> <ul style="list-style-type: none"> some of the crystals remained in the filtrate (after cooling and filtration) the solution was not allowed to cool for long enough (for complete crystallisation) magnesium carbonate is impure 	<p>IGNORE references to side reactions</p> <p>REJECT not all of the magnesium carbonate reacted</p>	1
Total			7

Question number	Answer	Notes	Marks
8 (a)	<p>M1 rate of forward reaction = rate of backward reaction</p> <p>M2 concentrations of reactants/products remain/stay constant</p>	<p>IGNORE forward reaction = backward reaction</p> <p>ACCEPT amounts/masses for concentrations ACCEPT do not change for remain constant ALLOW colour remains constant ALLOW pressure remains constant IGNORE concentrations/amounts of reactants and products are the same/are equal</p>	2

Question number	Answer	Notes	Marks
8 (b) (i)	 <p data-bbox="231 795 845 862">cross drawn on either curve just where it becomes horizontal</p> 		1
(ii)	<p data-bbox="231 1366 885 1444">M1 curve starts at 0,0 and has the general shape shown</p> <p data-bbox="231 1467 885 1579">M2 curve becomes horizontal at approximately the same time as the original curve</p> <p data-bbox="231 1612 885 1657">M3 curve finishes below original curve</p>	<p data-bbox="941 1612 1396 1724">M2 and M3 DEP on M1 or near miss (eg curve does not start at exactly 0,0)</p>	3

Question number	Answer	Notes	Marks
8 (c)	<p>M1 equilibrium has shifted to the left / equilibrium has shifted to the NO₂ side / equilibrium has shifted to the reactants side</p> <p>OR</p> <p>more NO₂ has been produced / more reactants have been produced</p> <p>M2 (therefore) backward reaction is endothermic</p>	<p>IGNORE references to Le Chatelier's principle eg an increase in temperature favours the reaction that reduces the temperature</p> <p>ACCEPT (forward) reaction is exothermic</p>	2
		Total	8

